

414 Thermodynamics and Thermochemistry

9. Heat produced in calories by the combustion of one gram of carbon is called
 (a) Heat of combustion of carbon
 (b) Heat of formation of carbon
 (c) Calorific value of carbon
 (d) Heat of production of carbon
10. Conditions of standard state used in thermochemistry is
 (a) $0^\circ C$ and 1 atm (b) $20^\circ C$ and 1 atm
 (c) $25^\circ C$ and 1 atm (d) $0 K$ and 1 atm
11. The temperature of the system decreases in an [KCET 2005]
 (a) Adiabatic compression (b) Isothermal compression
 (c) Isothermal expansion (d) Adiabatic expansion
12. For the isothermal expansion of an ideal gas
 (a) E and H increases
 (b) E increases but H decreases
 (c) H increases but E decreases
 (d) E and H are unaltered
13. If a refrigerator's door is opened, then we get [CPMT 1980]
 (a) Room heated
 (b) Room cooled
 (c) More amount of heat is passed out
 (d) No effect on room
14. The cooling in refrigerator is due to [KCET 2005]
 (a) Reaction of the refrigerator gas
 (b) Expansion of ice
 (c) The expansion of the gas in the refrigerator
 (d) The work of the compressor
15. Point out the wrong statement in relation to enthalpy
 (a) It is a state function
 (b) It is an intensive property
 (c) It is independent of the path followed for the change
 (d) Its value depends upon the amount of substance in the system
16. Which of the following is zero for an isochoric process
 (a) dP (b) dV
 (c) dT (d) dE
17. Mark the *correct* statement [MP PET 1997]
 (a) For a chemical reaction to be feasible, ΔG should be zero
 (b) Entropy is a measure of order in a system
 (c) For a chemical reaction to be feasible, ΔG should be positive
 (d) The total energy of an isolated system is constant
18. In an isochoric process the increase in internal energy is [MP PMT 1997]
 (a) Equal to the heat absorbed
 (b) Equal to the heat evolved
 (c) Equal to the work done
 (d) Equal to the sum of the heat absorbed and work done
19. It is general principle that the less energy a system contains, it is [MH CET 1999]
 (a) Less stable (b) More stable
 (c) Unstable (d) More unstable
20. Internal energy is an example of [Pb. PMT 2000]
 (a) Path function (b) State function
 (c) Both (a) and (b) (d) None of these
21. The process, in which no heat enters or leaves the system, is termed as [Pb. PMT 1999; Kerala (Med.) 2002; J & K 2005]
 (a) Isochoric (b) Isobaric
 (c) Isothermal (d) Adiabatic
22. The intensive property among these quantities is [KCET 2002]
 (a) Enthalpy (b) Mass/volume
 (c) Mass (d) Volume
23. In thermodynamics which one of the following is not an intensive property [Kerala (Med.) 2002]
 (a) Pressure (b) Density
 (c) Volume (d) Temperature
24. If in a container neither mass and nor heat exchange occurs then it constitutes a [MP PMT 1993; AFMC 2003]
 (a) Closed system (b) Open system
 (c) Isolated system (d) Imaginary system
25. Which of the following is not a state function [DCE 2002]
 (a) ΔS (b) ΔG
 (c) ΔH (d) ΔQ
26. Which of the following is true for an adiabatic process [DCE 2002; MP PET 1995, 98, 2004; CPMT 1990; MP PMT 1998, 2002]
 (a) $\Delta H = 0$ (b) $\Delta W = 0$
 (c) $\Delta Q = 0$ (d) $\Delta V = 0$
27. Which of the following is not a state function [MH CET 2004; AIIMS 2001]
 (a) Internal energy (b) Enthalpy
 (c) Work (d) Entropy
28. Among them intensive property is [AFMC 2004]
 (a) Mass (b) Volume
 (c) Surface tension (d) Enthalpy

First law of thermodynamics and Hess law

1. The first law of thermodynamics is only
 (a) The law of conservation of energy
 (b) The law of conservation of mass
 (c) The law of conservation of momentum
 (d) Both (a) and (b)
2. A mixture of two moles of carbon monoxide and one mole of oxygen, in a closed vessel is ignited to convert the carbon monoxide to carbon dioxide. If ΔH is the enthalpy change and ΔE is the change in internal energy, then [KCET 2005]
 (a) $\Delta H > \Delta E$
 (b) $\Delta H < \Delta E$
 (c) $\Delta H = \Delta E$
 (d) The relationship depends on the capacity of the vessel
3. Which of the following is always negative for exothermic reaction? [BCECE 2005]
 (a) ΔH (b) ΔS
 (c) ΔG (d) None of these

4. The relation between ΔE and ΔH is
 [MP PET 1992; MP PMT 1996; MP PET/PMT 1998]
 (a) $\Delta H = \Delta E - P \Delta V$ (b) $\Delta H = \Delta E + P \Delta V$
 (c) $\Delta E = \Delta V + \Delta H$ (d) $\Delta E = \Delta H + P \Delta V$
5. At constant T and P , which one of the following statements is correct for the reaction,
 $CO(g) + \frac{1}{2} O_2(g) \rightarrow CO_2(g)$
 [AIIMS 1982, 83; KCET 1988; BHU 1995; MP PET 1997, 99]
 (a) ΔH is independent of the physical state of the reactants of that compound
 (b) $\Delta H > \Delta E$
 (c) $\Delta H < \Delta E$
 (d) $\Delta H = \Delta E$
6. For the reaction of one mole of zinc dust with one mole of sulphuric acid in a bomb calorimeter, ΔU and w correspond to
 [AIIMS 2005]
 (a) $\Delta U < 0, w = 0$ (b) $\Delta U = 0, w < 0$
 (c) $\Delta U > 0, w = 0$ (d) $\Delta U < 0, w > 0$
7. Which is not the correct relation between enthalpy (ΔH) and intrinsic energy (ΔE)
 [EAMCET 1992]
 (a) $\Delta H = \Delta E + P \Delta V$ (b) $\Delta H = \Delta E + nRT$
 (c) $\Delta H = \Delta E - P \Delta V$ (d) $\Delta E = \Delta H - P \Delta V$
8. The law of Lavoisier and Laplace illustrates
 [KCET 1989]
 (a) The principle of conservation of energy
 (b) Equivalence of mechanical and thermal energy
 (c) The principle of conservation of matter
 (d) Equivalence of mechanical and chemical energy
9. For the reaction $N_2 + 3H_2 \rightleftharpoons 2NH_3; \Delta H =$
 [Roorkee 2000; CBSE PMT 1991, 2002]
 (a) $\Delta E - RT$ (b) $\Delta E - 2RT$
 (c) $\Delta E + RT$ (d) $\Delta E + 2RT$
10. If ΔH is the change in enthalpy and ΔE the change in internal energy accompanying a gaseous reaction
 [KCET 1989; CBSE PMT 1990]
 (a) ΔH is always greater than ΔE
 (b) $\Delta H < \Delta E$ only if the number of moles of the products is greater than the number of the reactants
 (c) ΔH is always less than ΔE
 (d) $\Delta H < \Delta E$ only if the number of moles of the products is less than the number of moles of the reactants
11. "The resultant heat change in a reaction is the same whether it takes place in one or several stages." This statement is called [MP PMT/PET 1988; MP PMT 1989]
 (a) Lavoisier and Laplace law
 (b) Hess's law
 (c) Joule's law
 (d) Le-chatelier's principle
12. Hess's law of constant heat summation is based on
 [MP PET 2001]
 (a) $E = mc^2$
 (b) Conservation of mass
 (c) First law of thermodynamics
 (d) None of the above
13. Consider the reaction : $N_2 + 3H_2 \rightarrow 2NH_3$ carried out at constant temperature and pressure. If ΔH and ΔU are the enthalpy and internal energy changes for the reaction, which of the following expression is true [AIEEE 2005]
 (a) $\Delta H = 0$ (b) $\Delta H = \Delta U$
 (c) $\Delta H < \Delta U$ (d) $\Delta H > \Delta U$
14. Which of the following is the correct equation
 [CBSE PMT 1996]
 (a) $\Delta U = \Delta Q - W$ (b) $\Delta W = \Delta U + \Delta Q$
 (c) $\Delta U = \Delta W + \Delta Q$ (d) None of these
15. Hess law is applicable for the determination of heat of
 [AIIMS 1998; Pb. PET/PMT 1999]
 (a) Reaction (b) Formation
 (c) Transition (d) All of these
16. Enthalpy for the reaction $C + O_2 \rightarrow CO_2$ is
 [DPMT 1987, 91]
 (a) Positive (b) Negative
 (c) Zero (d) None
17. During an isothermal expansion of an ideal gas its
 [CBSE PMT 1991]
 (a) Internal energy increases
 (b) Enthalpy decreases
 (c) Enthalpy remains unaffected
 (d) Enthalpy reduces to zero
18. The work done in ergs for the reversible expansion of one mole of an ideal gas from a volume of 10 litres to 20 litres at $25^\circ C$ is [CMC Vellore 1991]
 (a) $2.303 \times 298 \times 0.082 \log 2$
 (b) $298 \times 10^7 \times 8.31 \times 2.303 \log 2$
 (c) $2.303 \times 298 \times 0.082 \log 0.5$
 (d) $8.31 \times 10^7 \times 298 - 2.303 \log 0.5$
 (e) $2.303 \times 298 \times 2 \log 2$
19. The law of conservation of energy states that [NCERT 1984]
 (a) The internal energy of a system is constant
 (b) The heat content of a system is constant
 (c) Energy is neither created nor destroyed
 (d) There is an equivalence between energy and mass

416 Thermodynamics and Thermochemistry

- 20.** The enthalpies of the elements in their standard states are assumed to be
 (a) Zero at 298 K
 (b) Unit at 298 K
 (c) Zero at all temperatures
 (d) Zero at 273 K
- 21.** The heat Q for a reaction at constant volume is equal to
 (a) $H_P - H_R$ (b) $H_R - H_P$
 (c) $E_P - E_R$ (d) None of these
- 22.** Which of the following expressions represents the first law of thermodynamics
 [MP PET 1996, 2000; AFMC 1997; BHU 1999; AMU 2000; KCET (Med.) 2000, 01; CBSE PMT 2000; MP PMT 2002]
 (a) $\Delta E = -q + W$ (b) $\Delta E = q - W$
 (c) $\Delta E = q + W$ (d) $\Delta E = -q - W$
- 23.** In a reversible isothermal process, the change in internal energy is
 (a) Zero (b) Positive
 (c) Negative (d) None of these
- 24.** A schematic plot of $\ln K_{eq}$ versus inverse of temperature for a reaction is shown below
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- The reaction must be [AIEEE 2005]
 (a) Exothermic
 (b) Endothermic
 (c) One with negligible enthalpy change
 (d) Highly spontaneous at ordinary temperature
- 25.** Hess law of heat summation includes [AFMC 1992]
 (a) Initial reactants only
 (b) Initial reactants and final products
 (c) Final products only
 (d) Intermediates only
- 26.** An ideal gas at constant temperature and pressure expands, then its [BHU 1998]
 (a) Internal energy remains same
 (b) Internal energy decreases
 (c) Internal energy increases
 (d) Entropy first increases and then decreases
- 27.** The internal energy of a substance [KCET 1998; AFMC 2001; AIIMS 2001]
 (a) Increases with increase in temperature
 (b) Decreases with increase in temperature
 (c) Can be calculated by the relation $E = mc^2$
 (d) Remains unaffected with change in temperature
- 28.** The relation between change in internal energy (ΔE), change in enthalpy (ΔH) and work done (W) is represented as [AFMC 1999]
 (a) $\Delta H = \Delta E + W$ (b) $W = \Delta E - \Delta H$
 (c) $\Delta E = W - \Delta H$ (d) $\Delta E = \Delta H + W$
- 29.** The enthalpy of neutralization of which of the following acids and bases is nearly -13.6 Kcal [Roorkee 1999]
 (a) HCN and NaOH (b) HCl and KOH
 (c) HCl and NaOH (d) HCl and NH_4OH
- 30.** Work done during isothermal expansion of one mole of an ideal gas from 10 atm to 1 atm at 300 K is (Gas constant = 2) [AIIMS 2000]
 (a) 938.8 cal. (b) 1138.8 cal.
 (c) 1381.8 cal. (d) 1581.8 cal.
- 31.** Joule-Thomson expansion is [JIPMER 2000]
 (a) Isobaric (b) Isoenthalpic
 (c) Isothermal (d) None of these
- 32.** In which of the following $\Delta E = \Delta H$ [MP PMT 2001]
 (a) $\text{N}_2\text{O}_4(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$
 (b) $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$
 (c) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 (d) $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{l})$
- 33.** In an adiabatic expansion of an ideal gas [KCET (Med.) 2001; MH CET 2000]
 (a) $W = -\Delta E$ (b) $W = \Delta E$
 (c) $\Delta E = 0$ (d) $W = 0$
- 34.** For the reaction
 $\text{CH}_3\text{COOH}(\text{l}) + 2\text{O}_2(\text{g}) \rightleftharpoons 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$ at 25°C and 1 atm . pressure, $\Delta H = -874 \text{ kJ}$. Then the change in internal energy (ΔE) is [Orissa JEE 2002]
 (a) -874 kJ (b) -871.53 kJ
 (c) -876.47 kJ (d) $+874 \text{ kJ}$
- 35.** In a closed insulated container, a liquid is stirred with a paddle to increase its temperature. In this process, which of the following is true [CBSE PMT 2002]
 (a) $\Delta E = W = Q = 0$ (b) $\Delta E \neq 0, Q = W = 0$
 (c) $\Delta E = W \neq 0, Q = 0$ (d) $\Delta E = Q \neq 0, W = 0$
- 36.** Hess law is based on [MH CET 2002]
 (a) Law of conservation of mass
 (b) Law of conservation of energy
 (c) Enthalpy is a state function
 (d) None of these
- 37.** For the reaction,
 $\text{C}_3\text{H}_8(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 3\text{CO}_2(\text{g}) + 4\text{H}_2\text{O}(\text{l})$
 at constant temperature, $\Delta H - \Delta E$ is [CBSE PMT 2003]
 (a) $-RT$ (b) $+RT$
 (c) $-3RT$ (d) $+3RT$
- 38.** According to Hess's law, the heat of reaction depends upon

- (a) Initial condition of reactants
 (b) Initial and final conditions of reactants
 (c) Intermediate path of the reaction
 (d) End conditions of reactants
39. The value of $\Delta H - \Delta E$ for the following reaction at $27^\circ C$ will be $2NH_3(g) \rightarrow N_2(g) + 3H_2(g)$ [Kerala (Med.) 2002]
 (a) $8.314 \times 273 \times (-2)$ (b) $8.314 \times 300 \times (-2)$
 (c) $8.314 \times 27 \times (-2)$ (d) $8.314 \times 300 \times (2)$
40. For which of the following $\Delta E = \Delta H$ [MP PET 2003]
 (a) $N_2O_4(g) \rightleftharpoons 2NO_2(g)$
 (b) $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$
 (c) $H_2(g) + Cl_2(g) \rightleftharpoons 2HCl(g)$
 (d) $H_2(g) + \frac{1}{2}O_2(g) \rightleftharpoons H_2O(l)$
41. One mole of an ideal gas is allowed to expand freely and adiabatically into vacuum until its volume has doubled. A statement which is not true concerning this expression is [Pb. PMT 1998]
 (a) $\Delta H = 0$ (b) $\Delta S = 0$
 (c) $\Delta E = 0$ (d) $W = 0$
42. At $27^\circ C$, one mole of an ideal gas is compressed isothermally and reversibly from a pressure of 2 atm to 10 atm. The values of ΔE and q are ($R = 2$) [BHU 2001]
 (a) 0, -965.84 cal
 (b) -965.84 cal , $+965.84 \text{ cal}$
 (c) $+865.58 \text{ cal}$, -865.58 cal
 (d) -865.58 cal , -865.58 cal
43. The work done by a system is 8 joule, when 40 joule heat is supplied to it. What is the increase in internal energy of system [BHU 2001]
 (a) 25 J (b) 30 J
 (c) 32 J (d) 28 J
44. A system absorb 600J of heat and work equivalent to 300J on its surroundings. The change in internal energy is [Pb. PMT 2004]
 (a) 300 J (b) 400 J
 (c) 500 J (d) 600 J
45. Work done during isothermal expansion of one mole of an ideal gas from 10 atm. to 1 atm at 300K is [BHU 2004]
 (a) 4938.8 J (b) 4138.8 J
 (c) 5744.1 J (d) 6257.2 J
46. If gas, at constant temperature and pressure expands then its [MH CET 2003]
 (a) Entropy increases and then decreases
 (b) Internal energy increases
 (c) Internal energy remains the same
 (d) Internal energy decreases
47. For the reaction $PCl_5(g) \rightarrow PCl_3(g) + Cl_2(g)$ [MH CET 2004]
 (a) $\Delta H = \Delta E$ (b) $\Delta H > \Delta E$
 (c) $\Delta H < \Delta E$ (d) None of these
48. Enthalpy (H) is equal to [MH CET 2004]
 (a) Internal energy (E)
 (b) Product of pressure (P) and volume (V) of gas
 (c) Internal energy (E)+ PV
 (d) Work (W) done by a system
49. ΔE° of combustion of isobutylene is $-X \text{ kJ mol}^{-1}$. The value of ΔH° is [DCE 2004]
 (a) $= \Delta E^\circ$ (b) $> \Delta E^\circ$
 (c) $= 0$ (d) $< \Delta E^\circ$
50. One mole of an ideal gas is allowed to expand reversibly and adiabatically from a temperature of $27^\circ C$. If the work done during the process is 3 kJ, then final temperature of the gas is ($C_V = 20 \text{ J/K}$) [Pb. CET 2002]
 (a) 100 K (b) 150 K
 (c) 195 K (d) 255 K
51. Which of the following is correct regarding the internal energy of a substance [Pb. CET 2002]
 (a) It is a state function
 (b) It increases with increase in temperature
 (c) It can be calculated by experiment
 (d) It remains unaffected with change in temperature
52. For the gaseous reaction, $N_2O_4 \rightarrow 2NO_2$ [Pb. CET 2003]
 (a) $\Delta H < \Delta E$ (b) $\Delta H = \Delta E$
 (c) $\Delta H = 0$ (d) $\Delta H > \Delta E$
53. $2C + O_2 \rightarrow 2CO$; $\Delta H = -220 \text{ kJ}$
 Which of the following statement is correct for this reaction [BVP 2004]
 (a) Heat of combustion of carbon is 110 kJ
 (b) Reaction is exothermic
 (c) Reaction needs no initiation
 (d) All of these are correct
54. An ideal gas expands in volume from $1 \times 10^{-3} \text{ m}^3$ to $1 \times 10^{-2} \text{ m}^3$ at 300 K against a constant pressure of $1 \times 10^5 \text{ Nm}^{-2}$. The work done is [AIEEE 2004]
 (a) 270 kJ (b) -900 kJ
 (c) -900 J (d) 900 kJ
55. Internal energy is [AFMC 2004]
 (a) Partly potential and partly kinetic
 (b) Totally kinetic
 (c) Totally potential
 (d) None of these
56. For which of the reaction $\Delta H = \Delta E$ [MP PET 2004]

418 Thermodynamics and Thermochemistry

- (a) $N_2 + 3H_2 \rightleftharpoons 2NH_3$ (b) $N_2 + O_2 \rightarrow 2NO$
 (c) $2Na + Cl_2 \rightarrow 2NaCl$ (d) $PCl_5 \rightarrow PCl_3 + Cl_2$

57. The work done during the expansion of a gas from a volume of $4dm^3$ to $6dm^3$ against a constant external pressure of $3atm$ is ($1Latm = 101.32J$) [CBSE PMT 2004]

- (a) $+304J$ (b) $-304J$
 (c) $-6J$ (d) $-608J$

58. For the reaction, $A_{(s)} + 2B_{(g)} \rightarrow 4C_{(s)} + D_{(l)}$. ΔH and ΔU are related as [Orissa JEE 2004]

- (a) $\Delta H = \Delta U$ (b) $\Delta H = \Delta U + 3RT$
 (c) $\Delta H = \Delta U + RT$ (d) $\Delta H = \Delta U - 3RT$

IInd & IIIrd Law of thermodynamics and Entropy

1. For a reversible spontaneous change ΔS is

- (a) $\frac{\Delta E}{T}$ (b) $\frac{P\Delta V}{T}$
 (c) $\frac{q}{T}$ (d) $RT \log K$

2. When disorder of a system increases, the change is said to be

- (a) Exothermic (b) Non-spontaneous
 (c) Endothermic (d) Spontaneous

3. The spontaneous flow of heat is always

- (a) From low to high pressure
 (b) From high to high pressure
 (c) Unidirectional from lower temperature to higher temperature
 (d) Unidirectional from the higher to lower temperature

4. Mixing of non-reacting gases is generally accompanied by

- (a) Decrease in entropy
 (b) Increase in entropy
 (c) Change in enthalpy
 (d) Change in free energy

5. An irreversible process occurring isothermally in an isolated system leads to

- (a) Zero entropy
 (b) An increase in the total entropy of the system
 (c) A decrease in the total entropy of the system
 (d) None of these

6. The entropy values (in $JK^{-1}mol^{-1}$) of $H_{2(g)} = 130.6$, $Cl_{2(g)} = 223.0$ and $HCl_{(g)} = 186.7$ at $298K$ and $1atm$ pressure. Then entropy change for the reaction



- (a) $+540.3$ (b) $+727.3$
 (c) -166.9 (d) $+19.8$

7. Which of the following is the least random state of water

- (a) Ice
 (b) Liquid water
 (c) Steam
 (d) All present same random state

8. Which one of the following process is non-spontaneous

- (a) Dissolution of $CuSO_4$ in water
 (b) Reaction between H_2 and O_2 to form water

(c) Water flowing down hill

(d) Flow of electric current from low potential to high potential

9. Which of the following is zero during adiabatic expansion of the gas

- (a) ΔT (b) ΔS
 (c) ΔE (d) None of these

10. The entropy of crystalline substances at absolute zero going by the third law of thermodynamics should be taken as [J & K 2005]

- (a) 100
 (b) 50
 (c) Zero
 (d) Different for different substances

11. In which state, the matter have highest entropy

- (a) Solid (b) Liquid
 (c) Gas (d) Equal in all

12. Which of the following pairs of a chemical reaction is certain to result in spontaneous reaction? [CBSE PMT 2005]

- (a) Exothermic and decreasing disorder
 (b) Endothermic and increasing disorder
 (c) Exothermic and increasing disorder
 (d) Endothermic and decreasing disorder

13. When one mole of monoatomic ideal gas at $T K$ undergoes adiabatic change under a constant external pressure of $1atm$ changes volume from $1litre$ to $2litre$. The final temperature in Kelvin would be [IIT 2005]

- (a) $\frac{T}{2^{(2/3)}}$ (b) $T + \frac{2}{3 \times 0.0821}$
 (c) T (d) $T - \frac{2}{3 \times 0.0821}$

14. $9.0gm$ of H_2O is vaporised at $100^\circ C$ and $1atm$ pressure. If the latent heat of vaporisation of water is xJ/gm , then ΔS is given by

- (a) $\frac{x}{373}$ (b) $\frac{18x}{100}$
 (c) $\frac{18x}{373}$ (d) $\frac{1}{2} \times \frac{18x}{373}$

15. The ΔS for the vaporisation of $1mol$ of water is $88.3J/mole K$. The value of ΔS for the condensation of $1mol$ of vapour will be

- (a) $88.3J/mol K$ (b) $(88.3)^2 J/mol K$
 (c) $-88.3J/mol K$ (d) $\frac{1}{88.3} J/mol K$

16. For which reaction from the following, ΔS will be maximum [AIIMS 1982, 83, 94; AMU 2000]

- (a) $Ca(s) + \frac{1}{2}O_2(g) \rightarrow CaO(s)$
 (b) $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$
 (c) $C(s) + O_2(g) \rightarrow CO_2(g)$
 (d) $N_2(g) + O_2(g) \rightarrow 2NO(g)$

17. The occurrence of a reaction is impossible if
 [AIIMS 1982, 91; MP PET 1997; MP PMT 1999; DPMT 2002]
- (a) ΔH is +ve ; ΔS is also +ve but $\Delta H < T\Delta S$
 (b) ΔH is -ve ; ΔS is also -ve but $\Delta H > T\Delta S$
 (c) ΔH is -ve ; ΔS is +ve
 (d) ΔH is +ve ; ΔS is -ve
18. If the enthalpy of vaporization for water is $186.5 \text{ kJ mol}^{-1}$, the entropy of its vaporization will be
 [CPMT 1988; DCE 2000]
- (a) $0.5 \text{ JK}^{-1} \text{ mol}^{-1}$ (b) $1.0 \text{ JK}^{-1} \text{ mol}^{-1}$
 (c) $1.5 \text{ JK}^{-1} \text{ mol}^{-1}$ (d) $2.0 \text{ JK}^{-1} \text{ mol}^{-1}$
19. Which of the following statement is true. The entropy of the universe [MP PMT 1993; CET Pune 1998]
- (a) Increases and tends towards maximum value
 (b) Decreases and tends to be zero
 (c) Remains constant
 (d) Decreases and increases with a periodic rate
20. When enthalpy and entropy change for a chemical reaction are $-2.5 \times 10^3 \text{ cal}$ and 7.4 cal deg^{-1} respectively. Predict the reaction at 298 K is
 [AFMC 1998; MH CET 1999; CBSE PMT 2000]
- (a) Spontaneous (b) Reversible
 (c) Irreversible (d) Non-spontaneous
21. The total entropy change for a system and its surroundings increases, if the process is
- (a) Reversible (b) Irreversible
 (c) Exothermic (d) Endothermic
22. For chemical reactions, the calculation of change in entropy is normally done
- (a) At constant pressure
 (b) At constant temperature
 (c) At constant temperature and pressure both
 (d) At constant volume
23. When the value of entropy is greater, then the ability of work
- (a) Is maximum (b) Is minimum
 (c) Is medium (d) None of these
24. Which of the following is true for the reaction $H_2O(l) \rightleftharpoons H_2O(g)$ at 100°C and 1 atmosphere
 [KCET 1991; AIIMS 1996]
- (a) $\Delta E = 0$ (b) $\Delta H = 0$
 (c) $\Delta H = \Delta E$ (d) $\Delta H = T\Delta S$
25. The enthalpy change for the transition of liquid water to steam, $\Delta H_{\text{vap}} = 37.3 \text{ kJ mol}^{-1}$ at 373 K . The entropy change for the process is [MP PET 1994]
- (a) $111.9 \text{ J mol}^{-1} \text{ K}^{-1}$ (b) $37.3 \text{ J mol}^{-1} \text{ K}^{-1}$
 (c) $100 \text{ J mol}^{-1} \text{ K}^{-1}$ (d) $74.6 \text{ J mol}^{-1} \text{ K}^{-1}$
26. When a solid is converted into liquid, entropy [MP PET 1993, 95; MP PMT 1996]
- (a) Becomes zero (b) Remains the same
 (c) Decreases (d) Increases
27. In a spontaneous process, the entropy of the system and its surroundings [MP PET 1996; NCERT 1975]
- (a) Equals zero (b) Decreases
 (c) Increases (d) Remains constant
28. The positive value of ΔS indicates that [MP PMT 1997]
- (a) The system becomes less disordered
 (b) The system becomes more disordered
 (c) The system is in equilibrium position
 (d) The system tends to reach at equilibrium position
29. For reaction $2\text{Cl}(s) \rightarrow \text{Cl}_2(s)$, the signs of ΔH and ΔS respectively are [KCET 1996]
- (a) +, - (b) +, +
 (c) -, - (d) -, +
30. The enthalpy of water is 386 kJ . What is entropy of water [BHU 1997]
- (a) 0.5 kJ (b) 1.03 kJ
 (c) 1.5 kJ (d) 22.05 kJ
31. For which of the processes is ΔS negative [Pb. PMT 1998]
- (a) $\text{H}_2(g) \rightarrow 2\text{H}(g)$
 (b) $\text{N}_2(g) 1 \text{ atm} \rightarrow \text{N}_2(g) 8 \text{ atm}$
 (c) $2\text{SO}_3(g) \rightarrow 2\text{SO}_2(g) + \text{O}_2(g)$
 (d) $\text{C}_{(\text{diamond})} \rightarrow \text{C}_{(\text{graphite})}$
32. Identify the correct statement regarding entropy [CBSE PMT 1998; BHU 2001]
- (a) At 0°C , the entropy of a perfectly crystalline substance is taken to be zero
 (b) At absolute zero of temperature, the entropy of a perfectly crystalline substance is +ve
 (c) At absolute zero of temperature, the entropy of all crystalline substances is taken to be zero
 (d) At absolute zero of temperature, the entropy of a perfectly crystalline substance is taken to be zero
33. One mole of an ideal gas at 300 K is expanded isothermally from an initial volume of 1 litre to 10 litres. The change in energy for this process is ($R = 2 \text{ cal mol}^{-1} \text{ K}^{-1}$) [CBSE PMT 1998]
- (a) 163.7 cal (b) 850.2 cal
 (c) 1381.1 cal (d) Zero
34. If 900 J/g of heat is exchanged at boiling point of water, then what is increase in entropy [BHU 1998]
- (a) 43.4 J/mole (b) 87.2 J/mole
 (c) 900 J/mole (d) Zero
35. A container has hydrogen and oxygen mixture in ratio of 4 : 1 by weight, then [BHU 1998]
- (a) Internal energy of the mixture decreases
 (b) Internal energy of the mixture increases
 (c) Entropy of the mixture increases
 (d) Entropy of the mixture decreases
36. If 0.75 mole of an ideal gas expands isothermally at 27°C from 15 litres to 25 litres , the maximum work obtained is [AFMC 1999]
- (a) 8.40 J (b) 9.34 J
 (c) 10.86 J (d) 10.43 J

420 Thermodynamics and Thermochemistry

37. In which of the following conditions a chemical reaction can not occur [AMU 1999]
 (a) ΔH and ΔS increase and $T\Delta S > \Delta H$
 (b) ΔH and ΔS decrease and $\Delta H > T\Delta S$
 (c) ΔH increases and ΔS decreases
 (d) ΔH decreases and ΔS increases
38. An engine operating between $150^\circ C$ and $25^\circ C$ takes $500 J$ heat from a higher temperature reservoir if there are no frictional losses, then work done by engine is [MH CET 1999]
 (a) $147.7 J$ (b) $157.75 J$
 (c) $165.85 J$ (d) $169.95 J$
39. The entropy of a perfectly crystalline solid at absolute zero is [AMU 2000]
 (a) Positive (b) Negative
 (c) Zero (d) Not definite
40. Which of the following conditions will always lead to a non-spontaneous change [AMU 2000; KCET (Med.) 2000]
 (a) Positive ΔH and positive ΔS
 (b) Negative ΔH and negative ΔS
 (c) Positive ΔH and negative ΔS
 (d) Negative ΔH and positive ΔS
41. Following data is known about melting of a compound AB . $\Delta H = 9.2 kJ mol^{-1}$, $\Delta S = 0.008 kJ K^{-1} mol^{-1}$. Its melting point is [Pb. PMT 2000; AIIMS 2000]
 (a) $736 K$ (b) $1050 K$
 (c) $1150 K$ (d) $1150^\circ C$
42. The volume of a gas decreases from $500 cc$ to $300 cc$ when a sample of gas is compressed by an average pressure of $0.6 atm$. During this process $10 J$ of heat is liberated. The change in internal energy is [Pb. PMT 2000]
 (a) $-2.16 J$ (b) $12.156 J$
 (c) $2.16 J$ (d) $101.3 J$
43. The standard entropies of $CO_2(g)$, $C(s)$ and $O_2(g)$ are 213.5 , 5.690 and $205 JK^{-1}$ respectively. The standard entropy of formation of $CO_2(g)$ is [CPMT 2001]
 (a) $1.86 JK^{-1}$ (b) $1.96 JK^{-1}$
 (c) $2.81 JK^{-1}$ (d) $2.86 JK^{-1}$
44. Which of the following endothermic processes are spontaneous [CBSE PMT 2001]
 (a) Melting of ice (b) Evaporation of water
 (c) Heat of combustion (d) Both (a) and (b)
45. For the reaction $H_2O(s) \rightleftharpoons H_2O(l)$ at $0^\circ C$ and normal pressure [KCET 2001]
 (a) $\Delta H > T\Delta S$ (b) $\Delta H = T\Delta S$
 (c) $\Delta H = \Delta G$ (d) $\Delta H < T\Delta S$
46. One mole of an ideal gas is allowed to expand reversibly and adiabatically from a temperature of $27^\circ C$. If the work done during the process is $3 kJ$, the final temperature will be equal to ($C_v = 20 JK^{-1}$) [KCET 2000; AFMC 2000; AIIMS 2001]
 (a) $150 K$ (b) $100 K$
 (c) $26.85^\circ C$ (d) $295 K$
47. The entropy change, in the conversion of one mole of liquid water at $373 K$ to vapour at the same temperature is (Latent heat of vaporization of water, $\Delta H_{vap} = 2.257 kJ / g$) [AFMC 2002; KCET 2002]
 (a) $105.9 Jk^{-1}$ (b) $107.9 Jk^{-1}$
 (c) $108.9 Jk^{-1}$ (d) $109.9 Jk^{-1}$
48. A heat engine absorbs heat Q_1 at temperature T_1 and heat Q_2 at temperature T_2 . Work done by the engine is $(Q_1 + Q_2)$. This data [AIIEE 2002]
 (a) Violates Ist law of thermodynamics
 (b) Violates Ist law of thermodynamics if Q_1 is $-ve$
 (c) Violates Ist law of thermodynamics if Q_2 is $-ve$
 (d) Does not violate Ist law of thermodynamics
49. Equal volumes of monoatomic and diatomic gases at same initial temperature and pressure are mixed. The ratio of specific heats of the mixture (C_p / C_v) will be [AFMC 2002]
 (a) 1 (b) 2
 (c) 1.67 (d) 1.5
50. The law formulated by Dr. Nernst is [Kerala (Med.) 2002]
 (a) First law of thermodynamics
 (b) Second law of thermodynamics
 (c) Third law of thermodynamics
 (d) Both (a) and (b)
51. The unit of entropy is [CBSE PMT 2002]
 (a) $J mol^{-1}$ (b) $JK mol^{-1}$
 (c) $J mol^{-1} K^{-1}$ (d) $J^{-1} K^{-1} mol^1$
52. The entropy change involved in the conversion of 1 mole of liquid water at $373 K$ to vapour at the same temperature will be [$\Delta H_{vap} = 2.257 kJ / gm$] [MP PET 2002]
 (a) $0.119 kJ$ (b) $0.109 kJ$
 (c) $0.129 kJ$ (d) $0.120 kJ$
53. When a liquid boils, there is [JIPMER 2002]
 (a) An increase in entropy
 (b) A decrease in entropy
 (c) An increase in heat of vaporization
 (d) An increase in free energy
54. Which one of the following has ΔS° greater than zero [AIIMS 2003]
 (a) $CaO + CO_2(g) \rightleftharpoons CaCO_3(s)$
 (b) $NaCl(aq) \rightleftharpoons NaCl(s)$
 (c) $NaNO_3(s) \rightleftharpoons Na^+(aq) + NO_3^-(aq)$
 (d) $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$
55. The molar heat capacity of water at constant pressure is $75 JK^{-1} mol^{-1}$. When $1.0 kJ$ of heat is supplied to $100 g$ of water which is free to expand the increase in temperature of water is [CBSE PMT 2003]
 (a) $6.6 K$ (b) $1.2 K$
 (c) $2.4 K$ (d) $4.8 K$

56. The work done to contract a gas in a cylinder, is 462 joules. 128 joule energy is evolved in the process. What will be the internal energy change in the process [MP PMT 2003]
- (a) + 590 joules (b) - 334 joules
(c) + 334 joules (d) - 590 joules
57. Highest entropy is in [BVP 2003]
- (a) Hydrogen (b) Water
(c) Graphite (d) Mercury
58.
$$H^+(aq) + OH^-(aq) \longrightarrow H_2O(l)$$

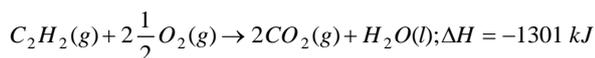
$$S^\circ(298 K) K^{-1} mol^{-1} \quad -10.7 \quad +70$$
 Standard entropy change for the above reaction is [DPMT 2004]
- (a) 60.3 JK⁻¹mol⁻¹ (b) 80.7 JK⁻¹mol⁻¹
(c) -70 JK⁻¹mol⁻¹ (d) +10.7 JK⁻¹mol⁻¹
59. One mole of water at 100°C is converted into steam at 100°C at a constant pressure of 1 atm. The change in entropy is [heat of vaporisation of water at 100°C = 540 cal / gm] [Pb. PMT 2004]
- (a) 8.74 (b) 18.76
(c) 24.06 (d) 26.06
60. For a carnot engine, the source is at 500 K and the sink at 300 K. What is efficiency of this engine [BHU 2004]
- (a) 0.2 (b) 0.4
(c) 0.6 (d) 0.3
61. Heat require to raise the temperature of 1 mol of a substance by 1°C is called [DCE 2002]
- (a) Specific heat (b) Molar heat capacity
(c) Water equivalent (d) Specific gravity
62. Maximum entropy will be in which of the following [DCE 2004]
- (a) Ice (b) Liquid water
(c) Snow (d) Water vapours
63. The standard entropies of CO₂(g), C(s) and O₂(g) are 213.5, 5.740 and 205 JK⁻¹ respectively. The standard entropy of formation of CO₂ is [Pb. CET 2001]
- (a) 2.76 JK⁻¹ (b) 2.12 JK⁻¹
(c) 1.12 JK⁻¹ (d) 1.40 JK⁻¹
64. Which law of thermodyanamics help in calculating entropy at different temperatures [Pb. CET 2003]
- (a) First law (b) Second law
(c) Third law (d) Zeroth law
65. Entropy is maximum in case of [BVP 2004]
- (a) Steam (b) Water at 0°C
(c) Water at 4°C (d) Ice
66. Considering entropy (S) as a thermodynamic parameter, the criterion for the spontaneity of any process is [CBSE PMT 2004]
- (a) ΔS_{system} > 0 only
(b) ΔS_{surroundin gs} > 0 only
(c) ΔS_{system} + ΔS_{surroundin gs} > 0
(d) ΔS_{system} - ΔS_{surroundin gs} > 0
67. A process is taking place at constant temperature and pressure. Then [KCET 2004]
- (a) ΔH = ΔE (b) ΔH = TΔS
(c) ΔH = 0 (d) ΔS = 0
68. The enthalpy of vapourization of a liquid is 30 kJ mol⁻¹ and entropy of vapourization is 75 J mol⁻¹K. The boiling point of the liquid at 1 atm is [IIT JEE (Screening) 2004]
- (a) 250K (b) 400K
(c) 450K (d) 600K

Heat of reaction

1. If $C + O_2 \rightarrow CO_2 + 94.2 \text{ kcal}$
 $H_2 + \frac{1}{2} O_2 \rightarrow H_2O + 68.3 \text{ kcal}$
 $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O + 210.8 \text{ kcal}$
 then the possible heat of methane will be [MP PET 1990]
- (a) 47.3 kcal (b) 20.0 kcal
(c) 45.9 kcal (d) -47.3 kcal
2. Molar heat of vaporisation of a liquid is 6kJ mol⁻¹. If the entropy change is 16J mol⁻¹ K⁻¹, the boiling point of the liquid is [KCET 2005]
- (a) 375°C (b) 375 K
(c) 273 K (d) 102°C
3. The enthalpy of fusion of ice per mole
- (a) 18 kJ (b) 8 kJ
(c) 80 kJ (d) 6 kJ
4. In which of the following neutralisation reactions, the heat of neutralisation will be highest [MP PMT 1989, 91; AIIMS 1999]
- (a) NH₄OH and CH₃COOH
(b) NH₄OH and HCl
(c) NaOH and CH₃COOH
(d) NaOH and HCl
5. If $S + O_2 \rightarrow SO_2; (\Delta H = -298.2)$
 $SO_2 + \frac{1}{2} O_2 \rightarrow SO_3; (\Delta H = -98.2)$
 $SO_3 + H_2O \rightarrow H_2SO_4; (\Delta H = -130.2)$
 $H_2 + \frac{1}{2} O_2 \rightarrow H_2O; (\Delta H = -287.3)$
 then the enthalpy of formation of H₂SO₄ at 298K will be [DPMT 1983; Orissa JEE 2005]
- (a) -433.7 kJ (b) -650.3 kJ
(c) +320.5 kJ (d) -813.9 kJ
6. From Kirchoff's equation which factor affects the heat of reaction [MP PMT 1990]
- (a) Pressure (b) Temperature
(c) Volume (d) Molecularity

422 Thermodynamics and Thermochemistry

7. The molar neutralization heat for KOH and HNO_3 as compared to molar neutralization heat of $NaOH$ and HCl
[MP PMT 1989]
- (a) Less (b) More
(c) Equal (d) Depends on pressure
8. An exothermic reaction is one in which the reacting substances
[CPMT 1974, 79; Bihar MEE 1982; KCET 1992; JIPMER 2001]
- (a) Have more energy than the products
(b) Have less energy than the products
(c) Are at a higher temperature than the product
(d) None of the above
9. The heat evolved in the combustion of benzene is given by
 $C_6H_6 + 7\frac{1}{2}O_2 \rightarrow 6CO_2(g) + 3H_2O(l); \Delta H = -3264.6 kJ$
Which of the following quantities of heat energy will be evolved when 39 g C_6H_6 are burnt
[NCERT 1978; MP PET 1990; JIPMER 2001]
- (a) 816.15 kJ (b) 1632.3 kJ
(c) 6528.2 kJ (d) 2448.45 kJ
10. Thermochemical reactions
 $C(\text{graphite}) + \frac{1}{2}O_2(g) \rightarrow CO(g); \Delta H = -110.5 kJ$
 $CO(g) + \frac{1}{2}O_2(g) \rightarrow CO_2(g); \Delta H = -283.2 kJ$
From the above reaction, the heat of reaction for $C(\text{graphite}) + O_2(g) \rightarrow CO_2(g)$ will be
[CPMT 1988; MP PMT 1989]
- (a) - 393.7 kJ (b) + 393.7 kJ
(c) - 172.7 kJ (d) + 172.7 kJ
11. The following is (are) endothermic reaction [IIT JEE 1999]
- (a) Combustion of methane
(b) Decomposition of water
(c) Dehydrogenation of ethane to ethylene
(d) Conversion of graphite to diamond
12. Evaporation of water is
[CPMT 1973; DPMT 1982; MP PMT 1989; MP PET 1999]
- (a) An endothermic change
(b) An exothermic change
(c) A process where no heat change occurs
(d) A process accompanied by chemical reaction
13. An exothermic reaction is one which
[NCERT 1977; MP PMT 1990]
- (a) Takes place only on heating
(b) Is accompanied by a flame
(c) Is accompanied by a absorption of heat
(d) Is accompanied by evolution of heat
14. An endothermic reaction is one in which
[MNR 1980; NCERT 1976]
- (a) Heat is converted into electricity
(b) Heat is absorbed
(c) Heat is evolved
(d) Heat is converted into mechanical work
15. Which of the following statement is correct
[NCERT 1978]
- (a) ΔH is positive for exothermic reaction
(b) ΔH is negative for endothermic reaction
(c) The heat of neutralization of strong acid and strong base is always the same
(d) The enthalpy of fusion is negative
16. The absolute enthalpy of neutralisation of the reaction $MgO(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2O(l)$ will be
[CBSE PMT 2005]
- (a) Less than $-57.33 kJ mol^{-1}$
(b) $-57.33 kJ mol^{-1}$
(c) Greater than $-57.33 kJ mol^{-1}$
(d) $57.33 kJ mol^{-1}$
17. ΔH_f° (298 K) of methanol is given by the chemical equation
[AIIMS 2005]
- (a) $CH_4(g) + 1/2O_2(g) \rightarrow CH_3OH(g)$
(b) $C(\text{graphite}) + 1/2O_2(g) + 2H_2(g) \rightarrow CH_3OH(l)$
(c) $C(\text{diamond}) + 1/2O_2(g) + 2H_2(g) \rightarrow CH_3OH(l)$
(d) $CO(g) + 2H_2(g) \rightarrow CH_3OH(l)$
18. If ΔH_f° for H_2O_2 and H_2O are $-188 kJ/mole$ and $-286 kJ/mole$. What will be the enthalpy change of the reaction $2H_2O_2(l) \rightarrow 2H_2O(l) + O_2(g)$
[MP PMT 1992]
- (a) $-196 kJ/mole$ (b) $146 kJ/mole$
(c) $-494 kJ/mole$ (d) $-98 kJ/mole$
19. The heat of transition (ΔH_t) of graphite into diamond would be, where
 $C(\text{graphite}) + O_2(g) \rightarrow CO_2(g); \Delta H = x kJ$
 $C(\text{diamond}) + O_2(g) \rightarrow CO_2(g); \Delta H = y kJ$ [Pb. PET 1985]
- (a) $(x + y)kJ mol^{-1}$ (b) $(x - y)kJ mol^{-1}$
(c) $(y - x)kJ mol^{-1}$ (d) None of these
20. Correct relationship between heat of fusion (ΔH_{fus}), heat of vaporization (ΔH_{vap}) and heat of sublimation (ΔH_{sub}) is
- (a) $\Delta H_{fus} = \Delta H_{vap} + \Delta H_{sub}$
(b) $\Delta H_{vap} = \Delta H_{fus} + \Delta H_{sub}$
(c) $\Delta H_{sub} = \Delta H_{vap} + \Delta H_{fus}$
(d) $\Delta H_{sub} = \Delta H_{vap} - \Delta H_{fus}$
21. Which of the following is an example of endothermic reaction
[MP PMT 1980]
- (a) $C_2H_2 + 2H_2 \rightarrow C_2H_6; \Delta E = -314.0 kJ$
(b) $C + O_2 \rightarrow CO_2; \Delta E = -393.5 kJ$
(c) $N_2 + O_2 \rightarrow 2NO; \Delta E = 180.5 kJ$
(d) $2H_2 + O_2 \rightarrow 2H_2O; \Delta E = 571.8 kJ$
22. If the enthalpy of B is greater than of A, the reaction $A \rightarrow B$ is
[MP PMT 1997]
- (a) Endothermic (b) Exothermic
(c) Instantaneous (d) Spontaneous
23. Given that
 $2C(s) + 2O_2(g) \rightarrow 2CO_2(g); \Delta H = -787 kJ$
 $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l); \Delta H = -286 kJ$



Heat of formation of acetylene is

[KCET 1989; Orissa JEE 2005]

- (a) -1802 kJ (b) $+1802 \text{ kJ}$
 (c) -800 kJ (d) $+228 \text{ kJ}$

24. Enthalpy change for reaction, $\frac{1}{2}H_2 + \frac{1}{2}Cl_2 \rightarrow HCl$, is called
 (a) Enthalpy of combination (b) Enthalpy of reaction
 (c) Enthalpy of formation (d) Enthalpy of fusion
25. The enthalpy of neutralization is about 57.3 kJ for the pair
 (a) HCl and NH_4OH
 (b) NH_4OH and HNO_3
 (c) HCl and $NaOH$
 (d) CH_3COOH and $NaOH$
26. Conversion of oxygen to ozone represented by the equation $3O_2 \rightarrow 2O_3$ is an endothermic reaction. Enthalpy change ΔH accompanying the reaction [KCET 1988, 92]
 (a) Is negative (b) Is positive
 (c) Is zero (d) Depends on temperature
27. The mutual heat of neutralisation of 40 gm of $NaOH$ and 60 gm CH_3COOH will be [MP PET/PMT 1988]
 (a) 56.1 kcal (b) Less than 56.1 kcal
 (c) More than 56.1 kcal (d) 13.7 kcal
28. The heat of formations of $CO(g)$ and $CO_2(g)$ are -26.4 kcal and -94.0 kcal respectively. The heat of combustion of carbon monoxide will be [MP PET/PMT 1988; EAMCET 1993]
 (a) $+26.4 \text{ kcal}$ (b) -67.6 kcal
 (c) -120.6 kcal (d) $+52.8 \text{ kcal}$
29. The enthalpy of combustion at $25^\circ C$ of H_2 , cyclohexane (C_6H_{12}) and cyclohexene (C_6H_{10}) are -241 , -3920 and -800 KJ / mole respectively. The heat of hydrogenation of cyclohexene is [BHU 2005]
 (a) -121 KJ / mole (b) $+121 \text{ KJ / mole}$
 (c) -242 KJ / mole (d) $+242 \text{ KJ / mole}$
30. Heat of neutralisation of NH_4OH and HCl is [EAMCET 1980; Roorkee 1990; MP PMT 1994]
 (a) 13.7 kcal (b) $>13.7 \text{ kcal}$
 (c) $<13.7 \text{ kcal}$ (d) None of the above
31. Heat of combustion ΔH of CH_4 , C_2H_6 , C_2H_4 and C_2H_2 gases are -212.8 , -373.0 , -337.0 and -310.5 kcal respectively at the same temperature. The best fuel among these gases is
 (a) CH_4 (b) C_2H_6
 (c) C_2H_4 (d) C_2H_2
32. Heat of formation of $CO_2(g)$, $H_2O(l)$ and $CH_4(g)$ are -94.0 , -68.4 and -17.9 kcal respectively. The heat of combustion of methane is
 (a) -212.9 kcal (b) -136.8 kcal
 (c) -304.3 kcal (d) -105.2 kcal

33. The heat of reaction does not depend upon
 (a) Temperature of the reaction
 (b) Physical state of reactants and products
 (c) Whether the reaction is carried out at constant pressure or at constant volume
 (d) The method by which the final products are obtained from the reactants
34. Heat of neutralisation of a strong acid by a strong base is a constant value because [KCET 1984]
 (a) Salt formed does not hydrolyse
 (b) Only H^+ and OH^- ions react in every case
 (c) The strong base and strong acid react completely
 (d) The strong base and strong acid react in aqueous solution
35. Heat of neutralisation of an acid by a base is highest when [KCET 1985]
 (a) Both the acid and base are weak
 (b) Both the acid and base are strong
 (c) The acid is strong and the base is weak
 (d) The acid is weak and the base is strong
36. The standard molar heat of formation of ethane, CO_2 and water (l) are respectively -21.1 , -94.1 and -68.3 kcal . The standard molar heat of combustion of ethane will be [IIT JEE 1986; DPMT 2005]
 (a) -372 kcal (b) 162 kcal
 (c) -240 kcal (d) 183.5 kcal
37. Which of the following reaction is endothermic [AFMC 1988]
 (a) $CaCO_3 \rightarrow CaO + CO_2$
 (b) $Fe + S \rightarrow FeS$
 (c) $NaOH + HCl \rightarrow NaCl + H_2O$
 (d) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$
38. In the reaction $CO_2(g) + H_2(g) \rightarrow CO(g) + H_2O(g); \Delta H = 80 \text{ kJ}$ ΔH is known as [CPMT 1987; KCET 1993; Bihar CEE 1995]
 (a) Heat of formation (b) Heat of combustion
 (c) Heat of neutralization (d) Heat of reaction
39. Heat of combustion of a substance [CPMT 1987, 96; AFMC 1992]
 (a) Is always positive
 (b) Is always negative
 (c) Is equal to heat of formation
 (d) Nothing can be said without reaction
40. The heat change for the reaction $H_2 + \frac{1}{2}O_2 \rightarrow H_2O$ is called
 (a) Heat of reaction (b) Heat of formation
 (c) Heat of vaporisation (d) None of the above
41. The heat change ΔH for the reaction $2CO + O_2 \rightarrow 2CO_2; \Delta H = -135 \text{ kcal}$ is called
 (a) Heat of formation (b) Heat of reaction
 (c) Heat of combustion (d) Heat of solution
42. The heats of combustion of rhombic and monoclinic sulphur are respectively 70960 and 71030 calories. What

424 Thermodynamics and Thermochemistry

will be the heat of conversion of rhombic sulphur to monoclinic

[MP PMT/PET 1988]

- (a) 70960 calories (b) 71030 calories
(c) -70 calories (d) +70 calories
43. The heat of formation of $H_2O(l)$ is -68.0 kcal , the heat of formation of $H_2O(g)$ is likely to be
(a) -68.0 kcal (b) -67.4 kcal
(c) 80.0 kcal (d) -58.3 kcal
44. Which of the following fuels will have the highest calorific value (kJ/kg) [MP PMT 1990]
(a) Charcoal (b) Kerosene
(c) Wood (d) Dung
45. The heat of combustion of carbon is -94 kcal at 1 atm pressure. The intrinsic energy of CO_2 is
(a) $+94 \text{ kcal}$ (b) -94 kcal
(c) $+47 \text{ kcal}$ (d) -47 kcal
46. The heat of neutralisation of a strong acid and a strong alkali is 57.0 kJ mol^{-1} . The heat released when 0.5 mole of HNO_3 solution is mixed with 0.2 mole of KOH is
[KCET 1991; AIIMS 2002; AMU (Engg.) 2002]
(a) 57.0 kJ (b) 11.4 kJ
(c) 28.5 kJ (d) 34.9 kJ
47. A solution of 500 ml of 0.2 M KOH and 500 ml of 0.2 M HCl is mixed and stirred; the rise in temperature is T_1 . The experiment is repeated using 250 ml each of solution, the temperature raised is T_2 . Which of the following is true [EAMCET 1987; MP PET 1994]
(a) $T_1 = T_2$ (b) $T_1 = 2T_2$
(c) $T_1 = 4T_2$ (d) $T_2 = 9T_1$
48. Enthalpy of a compound is equal to its [CMC Vellore 1991]
(a) Heat of combustion (b) Heat of formation
(c) Heat of reaction (d) Heat of solution
(e) Heat of dilution
49. Which is the best definition of "heat of neutralization" [CMC Vellore 1991]
(a) The heat set free when one gram molecule of a base is neutralized by one gram molecule of an acid in dilute solution at a stated temperature
(b) The heat absorbed when one gram molecule of an acid is neutralized by one gram molecule of a base in dilute solution at a stated temperature
(c) The heat set free or absorbed when one gram atom of an acid is neutralized by one gram atom of a base at a stated temperature
(d) The heat set free or absorbed when a normal solution containing one gram equivalent of an acid is neutralized by a normal solution containing one gram equivalent of a base at a stated temperature
(e) The heat set free when one gram equivalent of an acid is neutralized by one gram equivalent of a base in dilute solution at a stated temperature
50. The compound with negative heat of formation are known as [DPMT 1981]
(a) Endothermic compound
(b) Exothermic compound
(c) Heat of formation compound
(d) None of the above
51. If $H^+ + OH^- \rightarrow H_2O + 13.7 \text{ kcal}$, then the heat of neutralization for complete neutralization of one mole of H_2SO_4 by base will be [MP PMT 1990]
(a) 13.7 kcal (b) 27.4 kcal
(c) 6.85 kcal (d) 3.425 kcal
52. The lowest value of heat of neutralization is obtained for [KCET 1988; MP PMT 1990]
(a) $HCl + NaOH$
(b) $CH_3COOH + NH_4OH$
(c) $NH_4OH + HCl$
(d) $NaOH + CH_3COOH$
53. Heat of neutralisation for the given reaction $NaOH + HCl \rightarrow NaCl + H_2O$ is 57.1 kJ mol^{-1} . What will be the heat released when 0.25 mole of $NaOH$ is titrated against 0.25 mole of HCl [CPMT 1990]
(a) 22.5 kJ mol^{-1} (b) 57.1 kJ mol^{-1}
(c) 14.3 kJ mol^{-1} (d) 28.6 kJ mol^{-1}
54. Reaction, $H_2(g) + I_2(g) \rightarrow 2HI(g)$ $\Delta H = -12.40 \text{ kcal}$. According to this, the heat of formation of HI will be [MP PET 1990]
(a) 12.4 kcal (b) -12.4 kcal
(c) -6.20 kcal (d) 6.20 kcal
55. All reactions with chemical dissociation are [MP PMT 1990]
(a) Reversible
(b) Reversible and endothermic
(c) Exothermic
(d) Reversible or irreversible and endothermic or exothermic
56. The heat of formation is the change in enthalpy accompanying the formation of a substance from its elements at 298 K and 1 atm pressure. Since the enthalpies of elements are taken to be zero, the heat of formation (ΔH) of compounds [NCERT 1978; KCET 1993]
(a) Is always negative
(b) Is always positive
(c) May be negative or positive
(d) Is zero
57. The heat of combustion of $CH_{4(g)}$, $C_{(\text{graphite})}$ and $H_{2(g)}$ are 20 kcal , -40 kcal and -10 kcal respectively. The heat of formation of methane is [EAMCET 1998]
(a) -4.0 kcal (b) $+40 \text{ kcal}$
(c) -80 kcal (d) $+80 \text{ kcal}$
58. If a chemical reaction is accompanied by the evolution of heat, it is [BHU 1979]
(a) Catalytic (b) Photochemical
(c) Endothermic (d) Exothermic

59. A reaction that takes place with the absorption of energy is
[EAMCET 1977]
(a) Burning of a candle (b) Rusting of iron
(c) Electrolysis of water (d) Digestion of food
60. Which of the following is an endothermic reaction
[EAMCET 1980; MP PMT 1980; IIT JEE 1989; JIPMER 2002]
(a) $2H_2 + O_2 \rightarrow 2H_2O$
(b) $N_2 + O_2 \rightarrow 2NO$
(c) $2NaOH + H_2SO_4 \rightarrow Na_2SO_4 + 2H_2O$
(d) $3O_2 + C_2H_5OH \rightarrow 2CO_2 + 3H_2O$
61. One of the phenomena which cannot be described as combustion is
[EAMCET 1979]
(a) Oxidation of coal in air
(b) Burning of magnesium in nitrogen
(c) Reaction of antimony in chlorine
(d) Lighting of an electric lamp
62. Which of the following statements is correct about heat of combustion
[MADT Bihar 1982]
(a) It may be exothermic in some cases and endothermic in other cases
(b) It is applicable to gaseous substances only
(c) It is always an exothermic reaction
(d) Its value does not change with temperature
63. $H_2 + \frac{1}{2}O_2 \rightarrow H_2O; \Delta H = -68.39 \text{ kcal}$
 $K + H_2O + \text{Water} \rightarrow KOH(aq) + \frac{1}{2}H_2; \Delta H = -48 \text{ kcal}$
 $KOH + \text{Water} \rightarrow KOH(aq); \Delta H = -14 \text{ kcal}$
The heat of formation of KOH is (in kcal) [CPMT 1988]
(a) $-68.39 + 48 - 14$ (b) $-68.39 - 48 + 14$
(c) $68.39 - 48 + 14$ (d) $68.39 + 48 + 14$
64. Which of the following reaction can be used to define the heat of formation of $CO_2(g)$
[MP PMT 1989; MH CET 2001]
(a) $C(\text{graphite}) + O_2(g) = CO_2(g)$
(b) $CH_4(g) + 2O_2(g) = CO_2(g) + 2H_2O(l)$
(c) $CO(g) + \frac{1}{2}O_2(g) = CO_2(g)$
(d) $C_6H_6(l) + 7\frac{1}{2}O_2(g) = 6CO_2 + 3H_2O(l)$
65. The enthalpy of formation of ammonia is $-46.0 \text{ kJ mol}^{-1}$. The enthalpy change for the reaction $2NH_3(g) \rightarrow 2N_2(g) + 3H_2(g)$ is
(a) 46.0 kJ mol^{-1} (b) 92.0 kJ mol^{-1}
(c) $-23.0 \text{ kJ mol}^{-1}$ (d) $-92.0 \text{ kJ mol}^{-1}$
66. The heat of combustion of benzene determined in a bomb calorimeter is $-870 \text{ kcal mol}^{-1}$ at $25^\circ C$. The value of ΔE for the reaction is
(a) -1740 kcal (b) -870 kcal
(c) -872 kcal (d) -874 kcal
67. The formation of water from $H_2(g)$ and $O_2(g)$ is an exothermic reaction because [MP PMT/PET 1988]
(a) The chemical energy of $H_2(g)$ and $O_2(g)$ is more than that of water
(b) The chemical energy of $H_2(g)$ and $O_2(g)$ is less than that of water
(c) Not dependent on energy
(d) The temperature of $H_2(g)$ and $O_2(g)$ is more than that of water
68. In the reaction for the transition of carbon in the diamond form to carbon in the graphite form, ΔH is -453.5 cal . This points out that [BHU 1981; KCET 1986, 89]
(a) Graphite is chemically different from diamond
(b) Graphite is as stable as diamond
(c) Graphite is more stable than diamond
(d) Diamond is more stable than graphite
69. In the combustion of hydrocarbons, ΔH is [BHU 1981]
(a) Negative (b) Zero
(c) Positive (d) Undeterminate
70. The standard heats of formation in kcal mol^{-1} of $NO_2(g)$ and $N_2O_4(g)$ are 8.0 and 2.0 respectively. The heat of dimerization of NO_2 in kcal is
 $2NO_2(g) = N_2O_4(g)$ [NCERT 1983]
(a) 10.0 (b) -6.0
(c) -12.0 (d) -14.0
71. The difference between heats of reaction at constant pressure and at constant volume for the reaction $2C_6H_6(l) + 15O_2(g) \rightarrow 12CO_2(g) + 6H_2O(l)$ at $25^\circ C$ in kJ is
[IIT JEE 1991; Kerala PMT 2004]
(a) -7.43 (b) $+3.72$
(c) -3.72 (d) $+7.43$
72. $C(\text{diamond}) + C_2(g) \rightarrow CO_2(g); \Delta H = -395 \text{ kJ}$
 $C(\text{graphite}) + O_2(g) \rightarrow CO_2(g); \Delta H = -393.5 \text{ kJ}$
From the data, the ΔH when diamond is formed from graphite, is [CBSE PMT 1989; BHU 1987]
(a) -1.5 kJ (b) $+1.5 \text{ kJ}$
(c) $+3.0 \text{ kJ}$ (d) -3.0 kJ
73. Which of the following values of heat of formation indicates that the product is least stable [MP PMT 1991]
(a) -94 kcal (b) -231.6 kcal
(c) $+21.4 \text{ kcal}$ (d) $+64.8 \text{ kcal}$
74. Which of the following equations correctly represents the standard heat of formation (ΔH_f°) of methane
[IIT JEE (Screening) 1992]
(a) $C(\text{diamond}) + 2H_2(g) = CH_4(g)$
(b) $C(\text{graphite}) + 2H_2(g) = CH_4(l)$
(c) $C(\text{graphite}) + 2H_2(g) = CH_4(g)$

426 Thermodynamics and Thermochemistry

- (d) $C(\text{graphite}) + 4H = CH_4(g)$
75. In which of the following reactions does the heat change represent the heat of formation of water [EAMCET 1991]
- (a) $2H_2(g) + O_2(g) \rightarrow 2H_2O(l); \Delta H = -116 \text{ kcal}$
- (b) $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l); \Delta H = -58 \text{ kcal}$
- (c) $H^+(aq.) + OH^-(aq.) \rightarrow 2H_2O(l); \Delta H = -13.7 \text{ kcal}$
- (d) $C_2H_2(g) + 2\frac{1}{2}O_2(g) \rightarrow 2CO_2(g) + H_2O(l);$
 $\Delta H = -310 \text{ kcal}$
76. Values of heats of formation for SiO_2 and MgO are -48.4 and -34.7 kJ respectively. The heat of the reaction $2Mg + SiO_2 \rightarrow 2MgO + Si$ is [KCET 1990]
- (a) 21.16 kJ (b) -21.10 kJ
- (c) -13.62 kJ (d) 13.6 kJ
77. Based on the following thermochemical equations
- $H_2O(g) + C(s) \rightarrow CO(g) + H_2(g); \Delta H = 131 \text{ kJ}$
- $CO(g) + \frac{1}{2}O_2(g) \rightarrow CO_2(g); \Delta H = -282 \text{ kJ}$
- $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(g); \Delta H = -242 \text{ kJ}$
- $C(s) + O_2(g) \rightarrow CO_2(g); \Delta H = X \text{ kJ}$
- The value of X is [CBSE PMT 1992]
- (a) -393 kJ (b) -655 kJ
- (c) $+393 \text{ kJ}$ (d) $+655 \text{ kJ}$
78. Heat of transition is the heat evolved or absorbed when a substance is converted from [KCET 1984]
- (a) Solid to liquid
- (b) Solid to vapour
- (c) Liquid to vapour
- (d) One allotropic form to another allotropic form
79. For the allotropic change represented by equation $C(\text{diamond}) \rightarrow C(\text{graphite})$; the enthalpy change is $\Delta H = -1.89 \text{ kJ}$. If 6 g of diamond and 6 g of graphite are separately burnt to yield carbon dioxide, the heat liberated in the first case is [KCET 1988; DPMT 2000]
- (a) Less than in the second case by 1.89 kJ
- (b) More than in the second case by 1.89 kJ
- (c) Less than in the second case by 11.34 kJ
- (d) More than in the second case by 0.945 kJ
80. For an exothermic reaction [MP PET 1994; Manipal MEE 1995]
- (a) H of the products is less than H of the reactants
- (b) H of the products is more than H of the reactants
- (c) H of the products is equal to H of the reactants
- (d) ΔH is always positive
81. The energy evolved is highest for which of the following reactions [MP PET 1994]
- (a) $F + e^- \rightarrow F^{-1}$ (b) $Cl + e^- \rightarrow Cl^-$
- (c) $S + 2e^- \rightarrow S^{2-}$ (d) $O + 2e^- \rightarrow O^{2-}$
82. In the combustion of 2.0 gm of methane 25 kcal heat is liberated, heat of combustion of methane would be [MP PMT 1994]
- (a) 100 kcal (b) 200 kcal
- (c) 300 kcal (d) 400 kcal
83. Complete combustion of CH_4 gives [BHU 1995]
- (a) $CO_2 + H_2O$ (b) $CO_2 + H_2$
- (c) $COCl_2$ (d) $CO + CO_2 + H_2O$
84. If enthalpies of formation of $C_2H_4(g)$, $CO_2(g)$ and $H_2O(l)$ at $25^\circ C$ and 1 atm pressure be 52 , -394 and -286 kJ mol^{-1} respectively, the enthalpy of combustion of $C_2H_4(g)$ will be [CBSE PMT 1995; AIIMS 1998; Pb. PMT 1999]
- (a) $+1412 \text{ kJ mol}^{-1}$ (b) $-1412 \text{ kJ mol}^{-1}$
- (c) $+141.2 \text{ kJ mol}^{-1}$ (d) $-141.2 \text{ kJ mol}^{-1}$
85. Carbon and carbon monoxide burn in oxygen to form carbon dioxide according to the following reactions
- $C + O_2 \rightarrow CO_2; \Delta H = -394 \text{ kJ mol}^{-1}$
- $2CO + O_2 \rightarrow 2CO_2; \Delta H = -569 \text{ kJ mol}^{-1}$
- The heat of formation of 1 mol of carbon monoxide is thus [MP PET 1995]
- (a) $-219.0 \text{ kJ mol}^{-1}$ (b) $-109.5 \text{ kJ mol}^{-1}$
- (c) $-175.0 \text{ kJ mol}^{-1}$ (d) $-87.5 \text{ kJ mol}^{-1}$
86. Heat of formation in the reaction $H_2 + Cl_2 \rightarrow 2HCl + 44 \text{ kcal}$ is [MP PMT 1995]
- (a) 44 kcal (b) 44000 kcal
- (c) 22 kcal (d) 11 kcal
87. When water is added to quick lime, the reaction is [MP PMT 1995]
- (a) Exothermic (b) Endothermic
- (c) Explosive (d) None of these
88. In an exothermic reaction ΔH is [MP PMT 1995; CPMT 1999]
- (a) Positive (b) Negative
- (c) Zero (d) Both positive or negative
89. The heat change for the following reaction $C(s) + 2S(s) \rightarrow CS_2(l)$ is known as [MP PMT 1996]
- (a) Heat of vaporization (b) Heat of solution
- (c) Heat of fusion (d) Heat of formation
90. The enthalpy of combustion of benzene from the following data will be
- (i) $6C(s) + 3H_2(g) \rightarrow C_6H_6(l); \Delta H = +45.9 \text{ kJ}$
- (ii) $H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l); \Delta H = -285.9 \text{ kJ}$
- (iii) $C(s) + O_2(g) \rightarrow CO_2(g); \Delta H = -393.5 \text{ kJ}$

- (a) + 3172.8 kJ (b) - 1549.2 kJ [BHU 1997; DPMT 2002]
 (c) - 3172.8 kJ (d) - 3264.6 kJ
91. Enthalpy of formation of HF and HCl are -161 kJ and -92 kJ respectively. Which of the following statements is incorrect [KCET 2003]
 (a) HCl is more stable than HF
 (b) HF and HCl are exothermic compounds
 (c) The affinity of fluorine to hydrogen is greater than the affinity of chlorine to hydrogen
 (d) HF is more stable than HCl
92. The heat of reaction at constant pressure is given by [MP PMT 1997]
 (a) $E_p - E_R$ (b) $E_R - E_p$
 (c) $H_p - H_R$ (d) $H_R - H_p$
93. In an endothermic reaction, the value of ΔH is [CBSE PMT 1999; AIIMS 1998; BHU 2000; MP PMT 2000]
 (a) Zero (b) Positive
 (c) Negative (d) Constant
94. 2.2016 gm of acetaldehyde produced 13.95 kcal of heat on combustion in O_2 . Calculate the heat of combustion of CH_3CHO [Bihar CEE 1995]
 (a) 279 kcal (b) 972 kcal
 (c) 27.9 kcal (d) 2.79 kcal
95. $C + \frac{1}{2}O_2 \rightarrow CO; \Delta H = -42 \text{ kJ}$
 $CO + \frac{1}{2}O_2 \rightarrow CO_2; \Delta H = -24 \text{ kJ}$
 The heat of formation of CO_2 is [CPMT 1996]
 (a) - 16 kJ (b) + 66 kJ
 (c) + 16 kJ (d) - 66 kJ
96. Standard molar enthalpy of formation of CO_2 is equal to [IIT JEE 1997; BHU 2001]
 (a) Zero
 (b) The standard molar enthalpy of combustion of gaseous carbon
 (c) The sum of standard molar enthalpies of formation of CO and O_2
 (d) The standard molar enthalpy of combustion of carbon (graphite)
97. In the complete combustion of butanol $C_4H_9OH(l)$, if ΔH is enthalpy of combustion and ΔE is the heat of combustion at constant volume, then [EAMCET 1997]
 (a) $\Delta H < \Delta E$
 (b) $\Delta H = \Delta E$
 (c) $\Delta H > \Delta E$
 (d) $\Delta H, \Delta E$ relation cannot be predicted
98. $C + O_2 \rightarrow CO_2; \Delta H = X$
 $CO + \frac{1}{2}O_2 \rightarrow CO_2; \Delta H = Y$
 Then the heat of formation of CO is
 (a) $X - Y$ (b) $Y - 2X$
 (c) $X + Y$ (d) $2X - Y$
99. The values of heat of formation of SO_2 and SO_3 are -298.2 kJ and -98.2 kJ. The heat of reaction of the following reaction will be [BHU 1997; CBSE PMT 2000]
 $SO_2 + \frac{1}{2}O_2 \rightarrow SO_3$
 (a) -200 kJ (b) -356.2 kJ
 (c) +200 kJ (d) -396.2 kJ
100. Calculate the standard heat of formation of carbon disulphide (l), given that the standard heat of combustion of carbon (s), sulphur (s) and carbon disulphide (l) are -393.3, -293.72 and -1108.76 kJ mol⁻¹ respectively [Roorkee 1989; BHU 1997]
 (a) -128.02 kJ mol⁻¹ (b) +12.802 kJ mol⁻¹
 (c) +128.02 kJ mol⁻¹ (d) -12.802 kJ mol⁻¹
101. Fermentation is a reaction called [RPMT 1997]
 (a) Endothermic (b) Exothermic
 (c) Isotemperature (d) Reversible
102. 8 gm of CH_4 is completely burnt in air. The number of moles of water produced are [Orissa JEE 1997]
 (a) 0.5 (b) 1
 (c) 2 (d) 18
103. For the reaction $A(g) + 2B(g) \rightarrow 2C(g) + 3D(g)$, the value of ΔE at 27°C is 19.0 kcal. The value of ΔH for the reaction would be ($R = 2.0 \text{ cal K}^{-1} \text{ mol}^{-1}$) [Roorkee Qualifying 1998]
 (a) 20.8 kcal (b) 19.8 kcal
 (c) 18.8 kcal (d) 20.2 kcal
104. The enthalpy of solution of $BaCl_2(s)$ and $BaCl_2 \cdot 2H_2O(s)$ are -20.6 and 8.8 kJ mol⁻¹ respectively. The enthalpy change for the reaction $BaCl_2(s) + 2H_2O \rightarrow$ is [Roorkee Qualifying 1998]
 (a) 29.4 kJ (b) - 11.8 kJ
 (c) - 20.6 kJ (d) - 29.4 kJ
105. The enthalpy change of a reaction does not depend on [AIIMS 1997]
 (a) The state of reactants and products
 (b) Nature of reactants and products
 (c) Different intermediate reaction
 (d) Initial and final enthalpy change of a reaction
106. $S + \frac{3}{2}O_2 \rightarrow SO_3 + 2x \text{ kcal}$
 $SO_2 + \frac{1}{2}O_2 \rightarrow SO_3 + y \text{ kcal}$
 Find out the heat of formation of SO_2 [AIIMS 1997; CBSE PMT 1999; KCET 1999]
 (a) $2x - y$ (b) $2x + y$
 (c) $x + y$ (d) $2x / y$

428 Thermodynamics and Thermochemistry

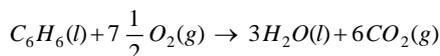
- 107.** The combustion enthalpies of carbon, hydrogen and methane are $-395.5 \text{ kJ mol}^{-1}$, $-285.8 \text{ kJ mol}^{-1}$ and $-890.4 \text{ kJ mol}^{-1}$ respectively at 25°C . The value of standard formation enthalpies of methane at that temperature is **[Pb. PMT 1998]**
- (a) $890.4 \text{ kJ mol}^{-1}$ (b) $-298.8 \text{ kJ mol}^{-1}$
 (c) $-74.7 \text{ kJ mol}^{-1}$ (d) $-107.7 \text{ kJ mol}^{-1}$
- 108.** Equal volumes of methanoic acid and sodium hydroxide are mixed. If x is the heat of formation of water, then heat evolved on neutralisation is **[BHU 1998]**
- (a) More than x (b) Equal to x
 (c) Twice of x (d) Less than x
- 109.** The neutralisation of a strong acid by a strong base liberates an amount of energy per mole of H^+ that **[BHU 1998]**
- (a) Depends upon which acid and base are involved
 (b) Depends upon the temperature at which the reaction takes place
 (c) Depends upon which catalyst is used
 (d) Is always the same
- 110.** When 0.5 g of sulphur is burnt to SO_2 , 4.6 kJ of heat is liberated. What is the enthalpy of formation of sulphur dioxide **[KCET 1998; AFMC 2001]**
- (a) $+147.2 \text{ kJ}$ (b) -147.2 kJ
 (c) -294.4 kJ (d) $+294.4 \text{ kJ}$
- 111.** Ozone is prepared by passing silent electric discharge through oxygen. In this reaction **[AFMC 1998]**
- (a) Energy is given out
 (b) Energy is absorbed
 (c) Oxygen is loaded with energy
 (d) Oxygen is dissociated into atoms
- 112.** Combustion of glucose takes place according to the equation,
 $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$, $\Delta H = -72 \text{ kcal}$. How much energy will be required for the production of 1.6 g of glucose (Molecular mass of glucose = 180 g) **[AFMC 1999]**
- (a) 0.064 kcal (b) 0.64 kcal
 (c) 6.4 kcal (d) 64 kcal
- 113.** Given that
 $\text{C}(s) + \text{O}_2(g) \rightarrow \text{CO}_2(g)$, $\Delta H = -394 \text{ kJ}$
 $2\text{H}_2(g) + \text{O}_2(g) \rightarrow 2\text{H}_2\text{O}(l)$, $\Delta H = 568 \text{ kJ}$
 $\text{CH}_4(g) + 2\text{O}_2(g) \rightarrow \text{CO}_2(g) + 2\text{H}_2\text{O}(l)$, $\Delta H = -892 \text{ kJ}$
 Heat of formation of CH_4 is
- (a) -70 kJ (b) 71.8 kJ
 (c) -244 kJ (d) $+782 \text{ kJ}$
- 114.** Calculate the heat of formation of $\text{PCl}_5(s)$ from the following data
 $2\text{P}(s) + 3\text{Cl}_2(g) \rightarrow 2\text{PCl}_3(l)$; $\Delta H = -151.8 \text{ kcal}$
 $\text{PCl}_3(l) + \text{Cl}_2(g) \rightarrow \text{PCl}_5(s)$; $\Delta H = -32.8 \text{ kcal}$ **[JIPMER 1999]**
- (a) -108.7 kcal (b) $+108.7 \text{ kcal}$
 (c) -184.6 kcal (d) $+184.6 \text{ kcal}$
- 115.** When 50 cm^3 of a strong acid is added to 50 cm^3 of an alkali, the temperature rises by 5°C . If 250 cm^3 of each liquid are mixed, the temperature rise would be **[KCET 1999]**
- (a) 5°C (b) 10°C
 (c) 25°C (d) 20°C
- 116.** $\text{H}_2(g) + \frac{1}{2}\text{O}_2(g) \rightarrow \text{H}_2\text{O}(l)$;
 ΔH at $298 \text{ K} = -285.8 \text{ kJ}$
 The molar enthalpy of vaporisation of water at 1 atm and 25°C is 44 kJ . The standard enthalpy of formation of 1 mole of water vapour at 25°C is **[KCET 1999]**
- (a) -241.8 kJ (b) 241.8 kJ
 (c) 329.8 kJ (d) -329.8 kJ
- 117.** When 4 g of iron is burnt to ferric oxide at constant pressure, 29.28 kJ of heat is evolved. What is the enthalpy of formation of ferric oxide (At. Wt. of $\text{Fe} = 56$) **[AIIMS 1999]**
- (a) -81.98 kJ (b) -819.8 kJ
 (c) -40.99 kJ (d) $+819.8 \text{ kJ}$
- 118.** When a strong acid, strong base or their salt are dissolved in water, they are completely ionised. If a strong acid is added to a strong base, H^+ ions from the former combine with OH^- ions of the latter forming water. The formation of each water molecule liberates a certain quantity of energy and the reaction is exothermic. The heat liberated when one mole of water is formed by combining hydrochloric acid and sodium hydroxide is 13.7 kcal . The heat liberated when one mole of water is formed by combining sulphuric acid and sodium hydroxide is **[AMU (Engg.) 1999]**
- (a) 25.5 kcal (b) 8.5 kcal
 (c) 13.7 kcal (d) 34 kcal
- 119.** 2.1 g of Fe combines with S evolving 3.77 kJ . The heat of formation of FeS in kJ/mol is **[AMU (Engg.) 1999]**
- (a) -1.79 (b) -100.5
 (c) -3.77 (d) None of these
- 120.** In the reaction: $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$, $\Delta H = 194 \text{ kJ}$. Heat of formation of HCl is **[KCET 2000; AFMC 2000; CPMT 2000]**
- (a) $+97 \text{ kJ}$ (b) $+194 \text{ kJ}$
 (c) -194 kJ (d) -97 kJ
- 121.** Enthalpy of neutralisation of acetic acid by NaOH is -50.6 kJ/mol . and the heat of neutralisation of a strong acid with a strong base is -55.9 kJ/mol . What is the value of ΔH for the ionisation of CH_3COOH **[AIIMS 2000]**
- (a) $+5.3 \text{ kJ/mol}$ (b) $+6.2 \text{ kJ/mol}$
 (c) $+8.2 \text{ kJ/mol}$ (d) $+9.3 \text{ kJ/mol}$
- 122.** Heat of neutralization of the acid-base reaction is 57.32 kJ for **[JIPMER 2000]**
- (a) $\text{HNO}_3 + \text{LiOH}$ (b) $\text{HCOOH} + \text{KOH}$
 (c) $\text{HCl} + \text{NH}_4\text{OH}$ (d) $\text{CH}_3\text{COOH} + \text{NaOH}$
- 123.** The heat change for the following reaction at 298°K and at constant pressure is $+7.3 \text{ kcal}$
 $\text{A}_2\text{B}(s) \rightarrow 2\text{A}(s) + 1/2 \text{B}_2(g)$, $\Delta H = +7.3 \text{ kcal}$

The heat change at constant volume would be

[DCE 2000]

- (a) 7.3 kcal (b) More than 7.3
(c) Zero (d) None of these

124. The heat evolved in the combustion of benzene is given by the equation



$$\Delta H = 781.0 \text{ kcal mol}^{-1}$$

Which of the following quantities of heat energy will be evolved when 39 g of benzene is burnt in an open container

[MP PET 2000]

- (a) 122.25 kcal mol⁻¹ (b) 390.5 kcal mol⁻¹
(c) 1562.0 kcal mol⁻¹ (d) 586.75 kcal mol⁻¹

125. $H_2(g) + Cl_2(g) \rightarrow 2HCl(g)$, $\Delta H = -44 \text{ kcal}$



For the reaction $Na(s) + \frac{1}{2}Cl_2(g) \rightarrow NaCl(s)$, $\Delta H =$

[KCET 2000]

- (a) -108 kcal (b) -196 kcal
(c) -98 kcal (d) 54 kcal

126. The ΔH_f° for $CO_2(g)$, $CO(g)$ and $H_2O(g)$ are -393.5, -110.5 and -241.8 kJ mol⁻¹ respectively. The standard enthalpy change (in kJ) for the reaction $CO_2(g) + H_2(g) \rightarrow CO(g) + H_2O(g)$ is [IIT JEE (Screening) 2000]

- (a) 524.1 (b) 41.2
(c) -262.5 (d) -41.2

127. Which of the following compounds will absorb the maximum quantity of heat when dissolved in the same amount of water? The heats of solution of these compounds at 25°C in kJ/mole of each solute is given in brackets [AMU (Engg.) 2000]

- (a) HNO_3 ($\Delta H = -33$) (b) KCl ($\Delta H = +17.64$)
(c) NH_4NO_3 ($\Delta H = +25.5$) (d) HCl ($\Delta H = -74.1$)

128. In the reaction $C + 2S \rightarrow CS_2 + \Delta H$, ΔH is the

[AMU (Engg.) 2000]

- (a) Heat of combustion (b) Heat of neutralisation
(c) Heat of solution (d) None of these

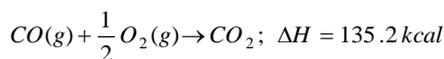
129. The heat of formation of methane $C(s) + 2H_2(g) \rightarrow CH_4(g)$ at constant pressure is 18500 cal at 25°C. The heat of reaction at constant volume would be [MH CET 2000]

- (a) 17904 cal (b) 18202 cal
(c) 18798 cal (d) 19096 cal

130. The enthalpy of combustion of $C_6H_6(l)$ is -3250 kJ. When 0.39 g of benzene is burnt excess of oxygen in an open vessel, the amount of heat evolved is [KCET 2000; AFMC 2000; DCE 2000]

- (a) 16.25 J (b) 16.25 kJ
(c) 32.5 J (d) 32.5 kJ

131. $C(s) + O_2(g) \rightarrow CO_2$; $\Delta H = -94 \text{ kcal}$



Then heat of formation of $CO(g)$ is [DCE 2001]

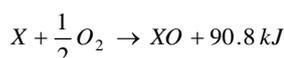
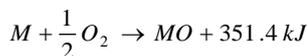
- (a) -26.4 kcal (b) 41.2 kcal
(c) 26.4 kcal (d) -229.2 kcal

132. The heat of neutralization of HCl and $NaOH$ is

[MP PET 2001]

- (a) Zero (b) -57.3 kJ
(c) +57.3 kJ (d) None of these

133. The following thermochemical reactions are given



It follows that the heat of reaction for the following process $M + XO \rightleftharpoons MO + X$ is given by

- (a) 422.2 kJ (b) 268.7 kJ
(c) -442.2 kJ (d) 260.6 kJ

134. If the heat of combustion of carbon monoxide at constant volume and at 17°C is -283.3 kJ, then its heat of combustion at constant pressure ($R = 8.314 \text{ J degree}^{-1} \text{ mol}^{-1}$) [CPMT 2001]

- (a) -284.5 kJ (b) 284.5 kJ
(c) 384.5 kJ (d) -384.5 kJ

135. Heat of formation of H_2O is -188 kJ/mole and H_2O_2 is -286 kJ/mole. The enthalpy change for the reaction $2H_2O_2 \rightarrow 2H_2O + O_2$ is [BHU 2001]

- (a) 196 kJ (b) -196 kJ
(c) 984 kJ (d) -984 kJ

136. The values of ΔH for the combustion of ethene and ethyne are -341.1 and -310.0 kcal respectively. Which of the following is a better fuel [BHU 2001]

- (a) C_2H_2 (b) C_2H_4
(c) Both of these (d) None of these

137. For exothermic reaction, the equilibrium constant

[JIPMER 2001]

- (a) Increases with increase of P
(b) Decreases with increase of P
(c) Increases with increase of temperature
(d) Decreases with increase of temperature

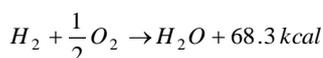
138. In order to decompose 9 g water 142.5 kJ heat is required. Hence the enthalpy of formation of water is [KCET 2001]

- (a) -142.5 kJ (b) +142.5 kJ
(c) -285 kJ (d) +285 kJ

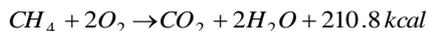
139. What is Δn for combustion of 1 mole of benzene, when both the reactants and the products are gas at 298 K [Pb. PMT 2001]

- (a) 0 (b) 3/2
(c) -3/2 (d) 1/2

140. If $C + O_2 \rightarrow CO_2 + 94.2 \text{ kcal}$



430 Thermodynamics and Thermochemistry



Then the possible heat of methane will be [UPSEAT 2001]

- (a) 47.3 kcal (b) 20.0 kcal
(c) 45.9 kcal (d) -47.3 kcal

141. Heat of neutralization of strong acid and weak base is

[UPSEAT 2001]

- (a) 57.1 kJ mol⁻¹
(b) 13.7 kJ mol⁻¹
(c) Less than 13.7 kcal mol⁻¹
(d) More than 13.7 kcal mol⁻¹

142. A system is changed from state A to state B by one path and from B to A another path. If E₁ and E₂ are the corresponding changes in internal energy, then

[Pb. PMT 2001]

- (a) E₁ + E₂ = -ve (b) E₁ + E₂ = +ve
(c) E₁ + E₂ = 0 (d) None of these

143. The heat evolved during the combination of 24 g C and 128 g S following the change is



- (a) 11 kcal (b) 32 kcal
(c) 44 kcal (d) 22 kcal

144. When the aqueous solution of 0.5 mole HNO₃ is mixed with the 0.3 mole of OH⁻ solution, then what will be the liberated heat (Enthalpy of neutralization is = 57.1 kJ)

[Kerala CET 2005]

- (a) 28.5 kJ (b) 17.1 kJ
(c) 45.7 kJ (d) 1.7 kJ
(e) 2.85 kJ

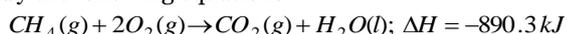
145. A cylinder of gas is assumed to contain 11.2 kg of butane (C₄H₁₀). If a normal family needs 20000 kJ of energy per day. The cylinder will last: (Given that ΔH for combustion of butane is -2658 kJ) [AMU (Engg.) 2002]

- (a) 20 days (b) 25 days
(c) 26 days (d) 24 days

146. Compounds with high heat of formation are less stable because [KCET 2002]

- (a) High temperature is required to synthesise them
(b) Molecules of such compounds are distorted
(c) It is difficult to synthesis them
(d) Energy rich state leads to instability

147. The heat evolved in the combustion of methane is given by the following equations:



How many grams of methane would be required to produce 445.15 kJ of heat of combustion

- (a) 4 g (b) 8 g
(c) 12 g (d) 16 g

148. Heats of combustion (ΔH°) for C(s), H₂(g) and CH₄(g) are -94, -68 and -213 kcal/mol. respectively. The value

of ΔH° for the reaction,
C(s) + 2H₂(g) → CH₄(g) is

[CBSE PMT 2002]

- (a) -85 kcal (b) -111 kcal
(c) -17 kcal (d) -170 kcal

149. If the value of ΔH in a reaction is positive, then the reaction is called [BHU 2002]

- (a) Exothermic (b) Endothermic
(c) Polymorphic (d) Polytypic

150. Enthalpy of neutralisation of NH₄OH and HCl, is numerically

[JIPMER 2002; Kurukshetra CEE 2002]

- (a) 57.1 kJ mol⁻¹ (b) < 57.1 kJ mol⁻¹
(c) > 57.1 kJ mol⁻¹ (d) Zero

151. The heat of neutralisation will be highest in [MP PMT 2002]

- (a) NH₄OH and CH₃COOH
(b) NH₄OH and HCl
(c) KOH and CH₃COOH
(d) KOH and HCl

152. If a mole of H₂ molecule is heated to high temperature the following reaction takes place [Kerala (Med.) 2002]

- (a) H₂(g) + 436 kJ = H(g) + H(g)
(b) H₂(g) + 820 kJ = 2H₂(g)
(c) 2H₂(g) + 436 J = 2H₂
(d) H₂ + H₂ = H⁺ + H⁺

153. Which of the following reactions is not exothermic

[MP PET 2002]

- (a) C(s) + O₂(g) → CO₂(g)
(b) C(s) + 2S(s) → CS₂(g)
(c) CH₄(g) + 2O₂(g) → CO₂(g) + 2H₂O(l)
(d) CO(g) + $\frac{1}{2}$ O₂(g) → CO₂(g)

154. On combustion, carbon forms two oxides CO and CO₂, Heat of formation of CO₂ gas is 94.3 kcal and that of CO is 26.0 kcal. Heat of combustion of carbon is [JIPMER 2002]

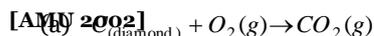
- (a) 26.0 kcal (b) -94.3 kcal
(c) 68.3 kcal (d) -120.3 kcal

155. 1 mole of conc. HCl requires X moles of dilute NaOH for neutralisation and 1 mole of concentrate H₂SO₄ requires Y moles of small dilute NaOH then which of the following reaction is true [MH CET 2002]

- (a) Y = $\frac{1}{2}$ X (b) X = $\frac{1}{2}$ Y
(c) X = 2Y (d) None of these

156. Which of the reaction defines ΔH_f

[IIT -JEE (Screening) 2003]



- (b) $\frac{1}{2}$ H₂(g) + $\frac{1}{2}$ F₂(g) → HF(g)
(c) N₂(g) + 3H₂(g) → 2NH₃(g)
(d) CO(g) + $\frac{1}{2}$ O₂(g) → CO₂(g)

157. One gram sample of NH₄NO₃ is decomposed in a bomb calorimeter. The temperature of the calorimeter increases

by 6.12 K the heat capacity of the system is 1.23 kJ/g/deg. What is the molar heat of decomposition for NH_4NO_3

[AIIMS 2003]

- (a) -7.53 kJ/mol (b) -398.1 kJ/mol
 (c) -16.1 kJ/mol (d) -602 kJ/mol

158. For which one of the following equations is $\Delta H_{\text{react}}^\circ$ equal to ΔH_f° for the product [CBSE PMT 2003]

- (a) $2CO(g) + O_2(g) \rightarrow 2CO_2(g)$
 (b) $N_2(g) + O_3(g) \rightarrow N_2O_3(g)$
 (c) $CH_4(g) + 2Cl_2(g) \rightarrow CH_2Cl_2(l) + 2HCl(g)$
 (d) $Xe(g) + 2F_2(g) \rightarrow XeF_4(g)$

159. Enthalpy change for a reaction does not depend upon

[AIEEE 2003]

- (a) The physical states of reactants and products
 (b) Use of different reactants for the same product
 (c) The nature of intermediate reaction steps
 (d) The differences in initial or final temperatures of involved substances

160. $C_{\text{dia}} + O_2 \rightarrow CO_2$; $\Delta H = -395.3 \text{ kJ/mole}$

$C_{\text{gr}} + O_2 \rightarrow CO_2$; $\Delta H = -393.4 \text{ kJ/mole}$

$C_{\text{gr}} \rightarrow C_{\text{dia}}$; $\Delta H =$

[BHU 2003]

- (a) -3.8 (b) -1.9
 (c) $+3.8$ (d) $+1.9$

161. What is the weight of oxygen that is required for the complete combustion of 2.8 kg of ethylene? [BVP 2003]

- (a) 9.6 kg (b) 96.0 kg
 (c) 6.4 kg (d) 2.8 kg

162. For complete neutralization of HCl with NaOH, the heat of neutralization is [MP PET 2003]

- (a) $+13.70 \text{ kJ mol}^{-1}$ (b) $-13.70 \text{ kJ mol}^{-1}$
 (c) $-57.32 \text{ kJ mol}^{-1}$ (d) $+57.32 \text{ kJ mol}^{-1}$

163. The heat of combustion of carbon to CO_2 is -393.5 kJ/mol . The heat released upon formation of 35.2 g of CO_2 from carbon and oxygen gas is [UPSEAT 2004]

- (a) $+315 \text{ kJ}$ (b) -31.5 kJ
 (c) -315 kJ (d) $+31.5 \text{ kJ}$

164. Which of the following equations correctly represents the standard heat of formation (ΔH_f°) of methane

[UPSEAT 2004]

- (a) $C(\text{diamond}) + 4H_{(g)} \rightarrow CH_{4(g)}$
 (b) $C(\text{diamond}) + 2H_{2(g)} \rightarrow CH_{4(g)}$
 (c) $C(\text{graphite}) + 2H_{2(g)} \rightarrow CH_{4(g)}$
 (d) $C(\text{graphite}) + 4H_{(g)} \rightarrow CH_{4(g)}$

165. If the heat of formation of CO_2 is -393 kJ . The amount of heat evolved in the formation of 0.156 kg of CO_2 is

[MH CET 2004]

- (a) -1357.9 kJ (b) -1275.9 kJ
 (c) -1572.0 kJ (d) -1165.5 kJ

166. Which of the following pairs has heat of neutralisation equal to 13.7 Kcals [DCE 2003]

- (a) HCl, NH_4OH (b) HNO_3, KOH
 (c) $NaOH, CH_3COOH$ (d) H_2SO_4, NH_4OH

167. The enthalpies of combustion of carbon and carbon monoxide are -393.5 and -283 kJ mol^{-1} respectively. The enthalpy of formation of carbon monoxide per mole is

[AIEEE 2004]

- (a) -676.5 kJ (b) 676.5 kJ
 (c) 110.5 kJ (d) -110.5 kJ

168. The enthalpy of combustion of methane at $25^\circ C$ is 890 kJ . The heat liberated when 3.2 g of methane is burnt in air is [KCET 2004]

- (a) 445 kJ (b) 278 kJ
 (c) -890 kJ (d) 178 kJ

169. If (i) $C + O_2 \rightarrow CO_2$, (ii) $C + 1/2O_2 \rightarrow CO$, (iii) $CO + 1/2O_2 \rightarrow CO_2$, the heats of reaction are $Q, -12, -10$ respectively. Then $Q =$ [Orissa JEE 2004]

- (a) -2 (b) 2
 (c) -22 (d) -16

170. How much energy is released when 6 mole of octane is burnt in air? Given ΔH_f° for $CO_2(g), H_2O(g)$ and $C_8H_{18}(l)$ respectively are $-490, -240$ and $+160 \text{ kJ/mol}$

[AIIMS 2004]

- (a) -6.2 kJ (b) -37.4 kJ
 (c) -35.5 kJ (d) -20.0 kJ

171. Given: $2Fe + \frac{3}{2}O_2 \rightarrow Fe_2O_3, \Delta H = -193.4 \text{ kJ}$;

$Mg + \frac{1}{2}O_2 \rightarrow MgO, \Delta H = -140.2 \text{ kJ}$

What is the ΔH of the reaction

$Mg + Fe_2O_3 \rightarrow 3MgO + 2Fe$

[Orissa JEE 2005]

- (a) -1802 kJ (b) $+1802 \text{ kJ}$
 (c) -800 kJ (d) $+228 \text{ kJ}$

172. The enthalpy change (ΔH) for the neutralisation of 1M HCl by caustic potash in dilute solution at 298 K is

[DPMT 2005]

- (a) 68 kJ (b) 65 kJ
 (c) 57.3 kJ (d) 50 kJ

432 Thermodynamics and Thermochemistry

1. If the bond dissociation energies of XY , X_2 and Y_2 (all diatomic molecules) are in the ratio of 1 : 1 : 0.5 and $\Delta_f H$ for the formation of XY is $-200 \text{ kJ mole}^{-1}$. The bond dissociation energy of X_2 will be **[AIEEE 2005]**
 - (a) 100 kJ mol^{-1}
 - (b) 800 kJ mol^{-1}
 - (c) 300 kJ mol^{-1}
 - (d) 400 kJ mol^{-1}
2. Energy required to dissociate 4 gm of gaseous hydrogen into free gaseous atoms is 208 kcal at 25°C . The bond energy of $H-H$ bond will be **[CPMT 1989; MP PET 2000; J & K 2005]**
 - (a) 104 kcal
 - (b) 10.4 kcal
 - (c) 1040 kcal
 - (d) 104 cal
3. The bond dissociation energy needed to form benzyl radical from toluene is than the formation of the methyl radical from methane **[IIT 1994]**
 - (a) Less
 - (b) Much
 - (c) Equal
 - (d) None of the above
4. Which one of the following bonds has the highest average bond energy (kcal/mol) **[CPMT 1994]**
 - (a) $S=O$
 - (b) $C\equiv C$
 - (c) $C\equiv N$
 - (d) $N\equiv N$
5. The bond dissociation energies of gaseous H_2 , Cl_2 and HCl are 104, 58 and 103 kcal respectively. The enthalpy of formation of HCl gas would be **[MP PET 1997; MP PMT 1999, 2001]**
 - (a) -44 kcal
 - (b) 44 kcal
 - (c) -22 kcal
 - (d) 22 kcal
6. The first ionization energy for Li is 5.4 eV and electron affinity of Cl is 3.61 eV. The ΔH (in kJ/mol) for the reaction $Li(g) + Cl(g) \rightarrow Li^+ + Cl^-$ is (if resulting ions do not combine with each other) ($1 \text{ eV} = 1.6 \times 10^{-19} \text{ J}$) **[MP PMT 2000, 03]**
 - (a) 70
 - (b) 100
 - (c) 170
 - (d) 270
7. Given that $C(g) + 4H(g) \rightarrow CH_4(g); \Delta H = -166 \text{ kJ}$
The bond energy $C-H$ will be **[AMU 2002]**
 - (a) 208 kJ/mole
 - (b) -41.6 kJ/mole
 - (c) 832 kJ/mole
 - (d) None of these
8. The $H-H$ bond energy is 430 kJ mol^{-1} and $Cl-Cl$ bond energy is 240 kJ mol^{-1} . ΔH for HCl is -90 kJ . The $H-Cl$ bond energy is about **[BVP 2003]**
 - (a) 180 kJ mol^{-1}
 - (b) 360 kJ mol^{-1}
 - (c) 213 kJ mol^{-1}
 - (d) 425 kJ mol^{-1}
9. If enthalpies of methane and ethane are respectively 320 and 360 calories then the bond energy of $C-C$ bond is **[UPSEAT 2003]**
 - (a) 80 calories
 - (b) 40 calories
 - (c) 60 calories
 - (d) 120 calories
10. If the bond energies of $H-H$, $Br-Br$ and HBr are 433, 192 and 364 kJ mol^{-1} respectively, the ΔH° for the reaction, $H_2(g) + Br_2(g) \rightarrow 2HBr(g)$ is **[CBSE PMT 2004]**
 - (a) $+261 \text{ kJ}$
 - (b) -103 kJ
 - (c) -261 kJ
 - (d) $+103 \text{ kJ}$

Free energy and Work function

1. The free energy change for a reversible reaction at equilibrium is **[NCERT 1984; Kurukshetra CEE 1998; AMU 1999]**
 - (a) Large positive
 - (b) Small negative
 - (c) Small positive
 - (d) 0
2. For a spontaneous change, free energy change ΔG is **[MNR 1983; BHU 1981, 95; AMU 1999; DCE 2000, 01; BHU 2000; MP PMT 2003]**
 - (a) Positive
 - (b) Negative
 - (c) Zero
 - (d) Can be positive or negative
3. A minus sign of the free energy change denotes that
 - (a) The reaction tends to proceed spontaneously
 - (b) The reaction is non-spontaneous
 - (c) The system is in equilibrium
 - (d) The reaction is very much unlikely
4. The relation between ΔG and ΔH is **[MP PMT 1994, 95, 97; AFMC 1997; Kurukshetra CEE 1998]**
 - (a) $\Delta H = \Delta G - T\Delta S$
 - (b) $\Delta G = \Delta H - T\Delta S$
 - (c) $T\Delta S - \Delta G = \Delta H$
 - (d) $\Delta H = T\Delta G + \Delta S$
5. At 300 K, the reactions which have following values of thermodynamic parameters occur spontaneously **[Roorkee 1999]**
 - (a) $\Delta G^\circ = -400 \text{ kJ mol}^{-1}$
 - (b) $\Delta H^\circ = 200 \text{ kJ mol}^{-1}$, $\Delta S^\circ = -4 \text{ JK}^{-1} \text{ mol}^{-1}$
 - (c) $\Delta H^\circ = -200 \text{ kJ mol}^{-1}$, $\Delta S^\circ = 4 \text{ JK}^{-1} \text{ mol}^{-1}$
 - (d) $\Delta H^\circ = 200 \text{ J mol}^{-1}$, $\Delta S^\circ = 40 \text{ JK}^{-1} \text{ mol}^{-1}$
6. The relation $\Delta G = \Delta H - T\Delta S$ was given by **[MP PMT 2000; KCET 2002]**
 - (a) Boltzmann
 - (b) Faraday
 - (c) Gibbs-Helmholtz
 - (d) Thomson
7. For precipitation reaction of Ag^+ ions with $NaCl$, which of the following statements is correct **[CPMT 1988]**
 - (a) ΔH for the reaction is zero
 - (b) ΔG for the reaction is zero
 - (c) ΔG for the reaction is negative
 - (d) $[\Delta G] = [\Delta H]$
8. At constant pressure and temperature, the direction to the result of any chemical reaction is where, there is less amount of
 - (a) Entropy
 - (b) Enthalpy
 - (c) Gibb's free energy
 - (d) None of the above

9. The dependence of Gibbs free energy on pressure for an isothermal process of an ideal gas is given by [MP PET 1996]
- (a) $\Delta G_T = nRT \ln \frac{P_2}{P_1}$ (b) $\Delta G_T = nRT \ln \frac{V_2}{V_1}$
 (c) $\Delta G_T = nRT \log \frac{P_1}{P_2}$ (d) $\Delta G_T = nRT \log \frac{V_2}{V_1}$
10. For the change $H_2O(l) \rightarrow H_2O(g)$; $P = 1 \text{ atm}$, $T = 373 \text{ K}$, the free energy change $\Delta G = 0$. This indicates that
- (a) $H_2O(l)$ is in equilibrium with $H_2O(g)$
 (b) Water boils spontaneously at 373 K
 (c) Water does not boil spontaneously at 373 K
 (d) Condensation of water vapour occurs spontaneously at 373 K
11. What is the free energy change ΔG when 1.0 mole of water at 100°C and 1 atm pressure is converted into steam at 100°C and 1 atm pressure [MP PET/PMT 1998]
- (a) 540 cal (b) -9800 cal
 (c) 9800 cal (d) 0 cal
12. ΔG° for the reaction $X + Y \rightleftharpoons Z$ is -4.606 kcal . The value of equilibrium constant of the reaction at 227°C is ($R = 2.0 \text{ cal} \cdot \text{mol}^{-1} \text{K}^{-1}$) [Roorkee 1999]
- (a) 100 (b) 10
 (c) 2 (d) 0.01
13. The standard enthalpy of the decomposition of N_2O_4 to NO_2 is 58.04 kJ and standard entropy of this reaction is 176.7 J/K . The standard free energy change for this reaction at 25°C is [AIIMS 1999]
- (a) 539 kJ (b) -539 kJ
 (c) -5.39 kJ (d) 5.39 kJ
14. Spontaneity of a chemical reaction is decided by the negative change in [MP PET 2001]
- (a) Internal energy (b) Enthalpy
 (c) Entropy (d) Free energy
15. For a reaction at 25°C enthalpy change and entropy changes are $-11.7 \times 10^3 \text{ J mol}^{-1}$ and $-105 \text{ J mol}^{-1} \text{K}^{-1}$ respectively. What is the Gibbs free energy [BHU 2001]
- (a) 15.05 kJ (b) 19.59 kJ
 (c) 2.55 kJ (d) 22.55 kJ
16. Born-Haber cycle is used to determine [UPSEAT 2001]
- (a) Crystal energy (b) Electron affinity
 (c) Lattice energy (d) All of these
17. Gibbs free energy G , enthalpy H and entropy S are interrelated as in [MP PMT 2002]
- (a) $G = H + TS$ (b) $G = H - TS$
 (c) $G - TS = H$ (d) $G = S = H$
18. The essential condition for the feasibility of a reaction is that [JIPMER 2002]
- (a) The reaction should be exothermic
 (b) The entropy of products must be larger than that of reactants
 (c) The reaction is to be accompanied with free energy decrease
 (d) The reaction has to possess high activation energy
19. The correct relationship between free energy change in a reaction and the corresponding equilibrium constant K_c is [AIIEE 2003]
- (a) $\Delta G = RT \ln K_c$ (b) $-\Delta G = RT \ln K_c$
 (c) $\Delta G^\circ = RT \ln K_c$ (d) $-\Delta G^\circ = RT \ln K_c$
20. In an irreversible process taking place at constant T and P and in which only pressure-volume work is being done, the change in Gibbs free energy (dG) and change in entropy (dS), satisfy the criteria [AIIEE 2003]
- (a) $(dS)_{V,E} < 0, (dG)_{T,P} < 0$
 (b) $(dS)_{V,E} > 0, (dG)_{T,P} < 0$
 (c) $(dS)_{V,E} = 0, (dG)_{T,P} = 0$
 (d) $(dS)_{V,E} = 0, (dG)_{T,P} > 0$
21. The densities of graphite and diamond at 298 K are 2.25 and 3.31 g cm^{-3} , respectively. If the standard free energy difference (ΔG°) is equal to 1895 J mol^{-1} , the pressure at which graphite will be transformed diamond at 298 K is [CBSE PMT 2003]
- (a) $9.92 \times 10^5 \text{ Pa}$ (b) $9.92 \times 10^8 \text{ Pa}$
 (c) $9.92 \times 10^7 \text{ Pa}$ (d) $9.92 \times 10^6 \text{ Pa}$
22. The free energy change for the following reactions are given below,
- $$C_2H_2(g) + \frac{5}{2} O_2(g) \rightarrow 2CO_2(g) + H_2O(l); \Delta G^\circ = -1234 \text{ kJ}$$
- $$C(s) + O_2(g) \rightarrow CO_2(g) \Delta G^\circ = -394 \text{ kJ}$$
- $$H_2(g) + \frac{1}{2} O_2(g) \rightarrow H_2O(l) \Delta G^\circ = -237 \text{ kJ}$$
- What is the standard free energy change for the reaction $H_2(g) + 2C(s) \rightarrow C_2H_2(g)$ [Kerala (Med.) 2002]
- (a) -209 kJ (b) -2259 kJ
 (c) $+2259 \text{ kJ}$ (d) 209 kJ
23. The equilibrium concentration of the species in the reaction $A + B \rightleftharpoons C + D$ are 3, 5, 10 and 15 mole L^{-1} respectively at 300 K the ΔG for the reaction is [Pb. PMT 2004]
- (a) 13.81 (b) -1381.8
 (c) -138.18 (d) 1391.6
24. Gibb's free energy (G) is defined as [Pb. CET 2001]
- (a) $\Delta G = \Delta H - T\Delta S$ (b) $\Delta G = \Delta H + \frac{T}{\Delta S}$
 (c) $\Delta H = \Delta G - T\Delta S$ (d) $\Delta G = \Delta H + T \cdot C_p$
25. Standard enthalpy and standard entropy changes for the oxidation of ammonia at 298 K are $-382.64 \text{ kJ mol}^{-1}$

434 Thermodynamics and Thermochemistry

and $-145.6 \text{ JK}^{-1}\text{mol}^{-1}$, respectively. Standard Gibbs energy change for the same reaction at 298 K is

[CBSE PMT 2004]

- (a) $-439.3 \text{ kJ mol}^{-1}$ (b) $-523.2 \text{ kJ mol}^{-1}$
 (c) $-221.1 \text{ kJ mol}^{-1}$ (d) $-339.3 \text{ kJ mol}^{-1}$

26. For spontaneity of a cell, which is correct

[Orissa JEE 2004]

- (a) $\Delta G = 0, \Delta E = 0$
 (b) $\Delta G = -ve, \Delta E = 0$
 (c) $\Delta G = +ve, \Delta E = +ve$
 (d) $\Delta G = -ve, \Delta E = +ve$

27. The free energy for a reaction having $\Delta H = 31400 \text{ cal}$; .

$\Delta S = 32 \text{ cal K}^{-1} \text{ mol}^{-1}$ at 1000°C is [Orissa JEE 2005]

- (a) -9336 cal (b) -7386 cal
 (c) -1936 cal (d) $+9336 \text{ cal}$

28. The ΔH and ΔS for a reaction at one atmospheric pressure are $+30.558 \text{ kJ}$ and 0.066 kJ K^{-1} respectively. The temperature at which the free energy change will be zero and below of this temperature the nature of reaction would be [Kerala CET 2005]

- (a) 483 K , spontaneous
 (b) 443 K , non-spontaneous
 (c) 443 K , spontaneous
 (d) 463 K , non-spontaneous
 (e) 463 K , spontaneous

Critical Thinking

Objective Questions

1. Adsorption of gases on solid surface is generally exothermic because [IIT JEE (Screening) 2004]

- (a) Enthalpy is positive (b) Entropy decreases
 (c) Entropy increases (d) Free energy increase

2. Two mole of an ideal gas is expanded isothermally and reversibly from 1 litre to 10 litre at 300 K . The enthalpy change (in kJ) for the process is

[IIT JEE (Screening) 2004]

- (a) 11.4 kJ (b) -11.4 kJ
 (c) 0 kJ (d) 4.8 kJ

3. Heat of neutralization of strong acid against strong base is constant and is equal to

[EAMCET 1980; AFMC 1989; DPMT 1991; MP PET 1999; BHU 1999; MP PMT 1995]

- (a) 13.7 kcal (b) 57 kJ
 (c) $5.7 \times 10^4 \text{ J}$ (d) All of the above

4. The value of heat generated when 36.5 gm HCl and 40 gm of NaOH reacts during neutralization

[NCERT 1984; CPMT 1993]

- (a) 76.5 kcal (b) 13.7 kcal

(c) More than 13.7 kcal (d) 108 kcal

5. When a gas undergoes adiabatic expansion, it gets cooled due to [DCE 2002]

- (a) Loss of kinetic energy
 (b) Fall in temperature
 (c) Decrease in velocity
 (d) Energy used in doing work

6. The heat content of a system is called

- (a) Internal energy (b) Entropy
 (c) Free energy (d) Enthalpy

7. To calculate the amount of work done in joules during reversible isothermal expansion of an ideal gas, the volume must be expressed in [KCET (Med.) 2001]

- (a) m^3 only (b) dm^3 only
 (c) cm^3 only (d) Any one of them

8. In an isobaric process, the ratio of heat supplied to the system (dQ) and work done by the system (dW) for diatomic gas is [AFMC 2002]

- (a) 1 : 1 (b) 7 : 2
 (c) 7 : 5 (d) 5 : 7

9. The enthalpy change for the reaction of 50.00 ml of ethylene with 50.00 ml of H_2 at 1.5 atm pressure is $\Delta H = -0.31 \text{ kJ}$. The value of ΔE will be [DPMT 2004]

- (a) -0.3024 kJ (b) 0.3024 kJ
 (c) 2.567 kJ (d) -0.0076 kJ

10. Enthalpy of solution of NaOH (solid) in water is $-41.6 \text{ kJ mol}^{-1}$. When NaOH is dissolved in water, the temperature of water [UPSEAT 2004]

- (a) Increase (b) Decreases
 (c) Does not change (d) Fluctuates indefinitely

11. In which of the following entropy decreases ?

[CPMT 1988, 94; MP PMT 2000]

- (a) Crystallization of sucrose from solution
 (b) Rusting of iron
 (c) Melting of ice
 (d) Vaporization of camphor

12. For conversion $\text{C}(\text{graphite}) \rightarrow \text{C}(\text{diamond})$ the ΔS is

[MP PMT 2001; MP PET 2003]

- (a) Zero (b) Positive
 (c) Negative (d) Unknown

13. For a reaction $\Delta H = 9.08 \text{ kJ mol}^{-1}$ and

$$\Delta S = 35.7 \text{ JK}^{-1}\text{mol}^{-1}$$

Which of the following statements is correct for the reaction

[AMU (Engg.) 2000]

- (a) Reversible and Isothermal
 (b) Reversible and Exothermic
 (c) Spontaneous and Endothermic
 (d) Spontaneous and Exothermic

14. For a reaction to occur spontaneously [CBSE PMT 1995]

- (a) $(\Delta H - T\Delta S)$ must be negative

- (b) $(\Delta H + T\Delta S)$ must be negative
 (c) ΔH must be negative
 (d) ΔS must be negative
15. The total amount of energy in the universe is fixed, but
 [AMU (Engg.) 1999]
 (a) Disorder is increasing (b) Lightning is increasing
 (c) Matter is increasing (d) Gravitation is decreasing
16. If for a given substance melting point is T_B and freezing point is T_A , then correct variation shown by graph between entropy change and temperature is [DCE 2001]
-
17. In endothermic reaction, the value of change in enthalpy (ΔH) is [Pb. CET 2001]
 (a) Positive (b) Negative
 (c) Zero (d) None of these
18. Which of the following would be expected to have the largest entropy per mole [MP PMT 2004]
 (a) $SO_2Cl_2(s)$ (b) $SO_2Cl_2(g)$
 (c) $SO_2Cl_2(l)$ (d) $SO_2(g)$
19. The enthalpies of formation of Al_2O_3 and Cr_2O_3 are $-1596 kJ$ and $-1134 kJ$ respectively. ΔH for the reaction $2Al + Cr_2O_3 \rightarrow 2Cr + Al_2O_3$ is [KCET 2003]
 (a) $-2730 kJ$ (b) $-462 kJ$
 (c) $-1365 kJ$ (d) $+2730 kJ$
20. Heat of reaction at constant volume is measured in the apparatus
 (a) Bomb calorimeter (b) Calorimeter
 (c) Pycnometer (d) Pyrometer
21. Which of the following gas has the highest heat of combustion
 (a) Methane (b) Ethane
 (c) Ethylene (d) Acetylene
22. The amount of heat measured for a reaction in a bomb calorimeter is [AIIMS 1991]
 (a) ΔG (b) ΔH
 (c) ΔE (d) $P\Delta V$
23. For an endothermic reaction where ΔH represents the enthalpy of the reaction in $kJ/mole$, the minimum value for the energy of activation will be [IIT JEE 1992; MP PET 1993]
 (a) Less than ΔH (b) Zero
 (c) More than ΔH (d) Equal to ΔH
24. The heat of neutralization of a strong acid by a strong base is nearly equal to [MP PET 1993; BCECE 2005]
 (a) $+57.32 J$ (b) $-57.32 kJ$
 (c) $+13.7 J$ (d) $-13.7 kJ$
25. A Beckmann thermometer is used to measure [Kurukshetra CEE 2002]
 (a) High temperature (b) Low temperature
 (c) Normal temperature (d) All temperature
26. The heat required to raise the temperature of a body by $1 K$ is called [AIEEE 2002]
 (a) Specific heat (b) Thermal capacity
 (c) Water equivalent (d) None of these
27. Mechanical work is specially important in system that contain [Kerala (Med.) 2002]
 (a) Solid-liquid (b) Liquid-liquid
 (c) Solid-solid (d) Amalgam
 (e) Gases
28. "The quantity of heat which must be supplied to decompose a compound into its element is equal to the heat evolved during the formation of that compound from the elements." This statement is known as
 (a) Hess's law
 (b) Joule's law
 (c) Le-chatelier's principle
 (d) Lavoiser and Laplace law
29. Hess law deals with [BHU 1979]
 (a) Changes in heat of reaction
 (b) Rate of reaction
 (c) Equilibrium constant
 (d) Influence of pressure on volume of a gas
30. For which one of the following reactions, ΔH is not equal to ΔE [IIT JEE 1995]
 (a) $H_{2(g)} + I_{2(g)} \rightleftharpoons 2HI_{(g)}$
 (b) $C_{(s)} + O_{2(g)} \rightleftharpoons CO_{2(g)}$
 (c) $N_{2(g)} + 3H_{2(g)} \rightleftharpoons 2NH_{3(g)}$
 (d) $HCl_{(aq)} + NaOH_{(aq)} \rightleftharpoons NaCl_{(aq)} + H_2O$
31. The heat liberated when $1.89 g$ of benzoic acid is burnt in a bomb calorimeter at $25^\circ C$ increases the temperature of $18.94 kg$ of water by $0.632^\circ C$. If the specific heat of water at $25^\circ C$ is $0.998 cal/g-deg$, the value of the heat combustion of benzoic acid is [CPMT 1999; BHU 2000]
 (a) $771.1 kcal$ (b) $871.2 kcal$
 (c) $881.1 kcal$ (d) $981.1 kcal$
32. For a hypothetical reaction $A \rightarrow B$, the activation energies for forward and backward reactions are $19 kJ/mole$ and $9 kJ/mole$ respectively. The heat of reaction is [CBSE PMT 2000]
 (a) $28 kJ$ (b) $19 kJ$
 (c) $10 kJ$ (d) $9 kJ$

436 Thermodynamics and Thermochemistry

33. The enthalpy of neutralization of HCN by $NaOH$ is $-12.13 \text{ kJ mol}^{-1}$. The enthalpy of ionisation of HCN will be
[BHU 1997]
(a) 4.519 kJ (b) 45.10 kJ
(c) 451.9 kJ (d) 45.19 kJ
34. In thermodynamics, a process is called reversible when
[AIIMS 2001]
(a) Surroundings and system change into each other
(b) There is no boundary between system and surroundings
(c) The surroundings are always in equilibrium with the system
(d) The system changes into the surroundings spontaneously
35. Which of the following unit represents largest amount of energy
[AMU 2001]
(a) Calorie (b) Joule
(c) Erg (d) Electron volt
36. Which of the following will have the highest bond energy
[MP PMT 1990]
(a) F_2 (b) Cl_2
(c) Br_2 (d) I_2
37. $C_{(\text{graphite})} + O_2(g) \rightarrow CO_2(g)$
 $\Delta H = -94.05 \text{ kcal mol}^{-1}$
 $C_{(\text{diamond})} + O_2(g) \rightarrow CO_2(g); \Delta H = -94.50 \text{ kcal mol}^{-1}$
therefore [DPMT 2001]
(a) $C_{(\text{graphite})} \rightarrow C_{(\text{diamond})}; \Delta H_{298 \text{ K}}^{\circ} = -450 \text{ cal mol}^{-1}$
(b) $C_{(\text{diamond})} \rightarrow C_{(\text{graphite})}; \Delta H_{298 \text{ K}}^{\circ} = +450 \text{ cal mol}^{-1}$
(c) Graphite is the stabler allotrope
(d) Diamond is harder than graphite
38. Enthalpy of formation of two compounds x and y are -84 kJ and -156 kJ respectively. Which of the following statements is correct
[CBSE PMT 2001]
(a) x is more stable than y
(b) x is less stable than y
(c) Both x and y are unstable
(d) x and y are endothermic compounds
39. For the process dry ice $\longrightarrow CO_2(g)$ [KCET 2000]
(a) ΔH is positive while $\Delta\rho$ is negative
(b) Both ΔH and $\Delta\rho$ are negative
(c) Both ΔH and $\Delta\rho$ are positive
(d) ΔH is negative while $\Delta\rho$ is positive
40. For melting of 3 moles of water at 0°C the ΔG° is
[MP PMT 2001]
(a) Zero (b) $+ve$
(c) $-ve$ (d) Unpredictable
41. The equilibrium constant of a reaction at 298 K is 5×10^{-3} and at 1000 K is 2×10^{-5} . What is the sign of ΔH for the reaction [Pb. CET 1998]
(a) $\Delta H = 0$ (b) ΔH is negative
(c) ΔH is positive (d) None of these
42. Which of the following has lowest fusion temperature
[MP PET 2003]
(a) Naphthalene (b) Diamond
(c) $NaCl$ (d) Mn
43. Consider the reactions
 $C(s) + 2H_2(g) \rightarrow CH_4(g), \Delta H = -x \text{ kcal}$
 $C(g) + 4H(g) \rightarrow CH_4(g), \Delta H = -x_1 \text{ kcal}$
 $CH_4(g) \rightarrow CH_3(g) + H(g), \Delta H = +y \text{ kcal}$
The bond energy of $C-H$ bond is [JIPMER 1997]
(a) $y \text{ kcal mol}^{-1}$ (b) $x_1 \text{ kcal mol}^{-1}$
(c) $x/4 \text{ kcal mol}^{-1}$ (d) $x_1/4 \text{ kcal mol}^{-1}$
44. Given the bond energies $N \equiv N$, $H-H$ and $N-H$ bonds are $945, 436$ and 391 kJ mole^{-1} respectively, the enthalpy of the following reaction $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$ is
[EAMCET 1992; JIPMER 1997]
(a) -93 kJ (b) 102 kJ
(c) 90 kJ (d) 105 kJ
45. The relation between ΔG and E for a cell is $\Delta G = -nFE$; the cell reaction will be spontaneous if [MP PET 1995]
(a) G is negative (b) G is positive
(c) E is negative (d) E is positive
46. Which of the following is not a correct statement?
[AMU (Engg.) 2002]
(a) When ΔG is negative, the process is spontaneous
(b) When ΔG is zero, the process is in a state of equilibrium
(c) When ΔG is positive, the process is non-spontaneous
(d) None of these
47. If at 298 K the bond energies of $C-H, C-C, C=C$ and $H-H$ bonds are respectively $414, 347, 615$ and 435 kJ mol^{-1} , the value of enthalpy change for the reaction $H_2C = CH_2(g) + H_2(g) \rightarrow H_3C-CH_3(g)$ at 298 K will be
[AIEEE 2003]
(a) $+250 \text{ kJ}$ (b) -250 kJ
(c) $+125 \text{ kJ}$ (d) -125 kJ
48. The equation
 $\frac{1}{2}H_2 + \frac{1}{2}Cl_2 \rightarrow HCl$ ($\Delta H_{298} = -22060 \text{ kcal}$) means
[CMC Vellore 1991]
(a) The heat absorbed when one gram molecule of HCl is formed from its elements at 25°C is 22.060 kcal
(b) The heat given out when one gram molecule of HCl is formed from its elements at 298 K is 22.060 kcal
(c) The heat absorbed when one atom of hydrogen reacts with one atom of chlorine to form one molecule of HCl at 25°C and one atmospheric pressure is 22.060 kcal

- (d) The heat absorbed when one gram equivalent of HCl is formed from its elements at 298 K is 22.060 kcal
- (e) The intrinsic heat of one molecule of HCl is 22.060 kcal more than the intrinsic heats of one atom of hydrogen and one atom of chlorine
49. The $H_2O(g)$ molecule dissociates as
- (i) $H_2O(g) \rightarrow H(g) + OH(g); \Delta H = 490\text{ kJ}$
- (ii) $OH(g) \rightarrow H(g) + O(g); \Delta H = 424\text{ kJ}$
- The average bond energy (in kJ) for water is
- (a) 490 (b) 424
(c) 914 (d) $914/2$
50. When 50 cm^3 of $0.2\text{ N } H_2SO_4$ is mixed with 50 cm^3 of 1 N KOH , the heat liberated is [KCET 2004]
- (a) 11.46 kJ (b) 57.3 kJ
(c) 573 kJ (d) 573 J
51. Following reaction occurring in an automobile $2C_8H_{18}(g) + 25O_2(g) \rightarrow 16CO_2(g) + 18H_2O(g)$. The sign of $\Delta H, \Delta S$ and ΔG would be [CBSE PMT 1994; KCET 1999]
- (a) +, -, + (b) -, +, -
(c) -, +, + (d) +, +, -
52. For the reaction
- $$H_2(g) + \frac{1}{2}O_2(g) \rightarrow H_2O(l), \Delta H = -285.8\text{ kJ mol}^{-1}$$
- $\Delta S = -0.163\text{ kJ mol}^{-1}K^{-1}$. What is the value of free energy change at 27° C for the reaction [KCET 1999]
- (a) $-236.9\text{ kJ mol}^{-1}$ (b) $-281.4\text{ kJ mol}^{-1}$
(c) $-334.7\text{ kJ mol}^{-1}$ (d) $+334.7\text{ kJ mol}^{-1}$
- Reason : The volume occupied by the molecules of an ideal gas is zero. [AIIMS 2002]
5. Assertion : Absolute values of internal energy of substance can not be determined.
- Reason : It is impossible to determine exact values of constituent energies of the substances. [AIIMS 2002]
6. Assertion : Mass and volume are extensive properties.
- Reason : Mass / volume is also an extensive parameter. [AIIMS 2002]
7. Assertion : Molar entropy of vaporization of water is different from ethanol.
- Reason : Water is more polar than ethanol.
8. Assertion : The increase in internal energy (ΔE) for the vaporization of one mole of water at 1 atm and 373 K is zero.
- Reason : For all isothermal processes $\Delta E = 0$. [AIIMS 2003]
9. Assertion : ΔH and ΔE are almost the same for the reaction. $N_2(g) + O_2(g) \rightleftharpoons 2NO(g)$.
- Reason : All reactants and products are gases. [AIIMS 2003]
10. Assertion : The enthalpies of neutralisation of strong acids and strong bases are always same.
- Reason : Neutralisation is heat of formation of water. [AIIMS 1996]
11. Assertion : Zeroth law can also be termed as law of thermal equilibrium.
- Reason : Two objects in thermal equilibrium with the third one, are in thermal equilibrium with each other.
12. Assertion : There is no reaction known for which ΔG is positive, yet it is spontaneous.
- Reason : For photochemical reactions ΔG is negative.
13. Assertion : Heat of neutralisation of perchloric acid, $HClO_4$ with $NaOH$ is same as is that of HCl with $NaOH$.
- Reason : Both HCl and $HClO_4$ are strong acids.
14. Assertion : Heat absorbed in a reaction at constant temperature and constant volume is $-\Delta G$.
- Reason : ΔG should be negative for the reaction to be spontaneous.
15. Assertion : T, P and V are state variables or state functions.
- Reason : Their values depends on the state of the system and how it is reached.
16. Assertion : Internal energy is an extensive property.
- Reason : Internal energy depends upon the amount of the system.
17. Assertion : For the combustion reactions, the value of ΔH is always negative.

Assertion & Reason

For AIIMS Aspirants

Read the assertion and reason carefully to mark the correct option out of the options given below :

- (a) If both assertion and reason are true and the reason is the correct explanation of the assertion.
- (b) If both assertion and reason are true but reason is not the correct explanation of the assertion.
- (c) If assertion is true but reason is false.
- (d) If the assertion and reason both are false.
- (e) If assertion is false but reason is true.

1. Assertion : The enthalpy of formation of gaseous oxygen molecules at 298 K and under pressure of 1 atm is zero.
- Reason : The entropy of formation of gaseous oxygen molecules under the same condition is zero. [AIIMS 1996]
2. Assertion : We feel cold on touching the ice.
- Reason : Ice is a solid form of water. [AIIMS 1999]
3. Assertion : Entropy of ice is less than water.
- Reason : Ice have cage like structure. [AIIMS 2000]
4. Assertion : The heat absorbed during the isothermal expansion of an ideal gas against vacuum is zero.

438 Thermodynamics and Thermochemistry

- Reason : The combustions reactions are always endothermic.
18. Assertion : For an isothermal reversible process $Q = -w$ i.e. work done by the system equals the heat absorbed by the system.
- Reason : Enthalpy change (ΔH) is zero for isothermal process.
19. Assertion : Enthalpy of formation of graphite is zero but of diamond is not zero.
- Reason : Enthalpy of formation of most stable allotrope is taken as zero.
20. Assertion : If a refrigerator's door is kept open room gets cooled.
- Reason : Material kept inside the refrigerator remains cool.
21. Assertion : Enthalpy and entropy of any elementary substance in the standard state are taken as zero.
- Reason : At zero degree absolute, the constituent particles become completely motionless.
22. Assertion : A process is called adiabatic if the system does not exchange heat with the surroundings.
- Reason : It does not involve increase or decrease in temperature of the system.

41	d	42	a	43	c	44	a	45	c
46	c	47	b	48	c	49	d	50	b
51	a	52	d	53	b	54	c	55	a
56	b	57	d	58	d				

IInd & IIIrd Law of thermodynamics and Entropy

1	c	2	d	3	d	4	b	5	b
6	d	7	a	8	d	9	b	10	c
11	c	12	c	13	d	14	d	15	c
16	b	17	d	18	a	19	a	20	a
21	d	22	c	23	a	24	d	25	c
26	d	27	c	28	b	29	c	30	b
31	b	32	d	33	d	34	a	35	c
36	b	37	c	38	a	39	c	40	c
41	c	42	c	43	c	44	d	45	b
46	a	47	c	48	d	49	d	50	c
51	c	52	b	53	a	54	c	55	c
56	c	57	a	58	b	59	d	60	b
61	b	62	d	63	a	64	c	65	a
66	c	67	b	68	b				

Answers

Basic concepts

1	b	2	d	3	b	4	c	5	c
6	d	7	d	8	c	9	c	10	c
11	d	12	d	13	a	14	c	15	b
16	b	17	d	18	a	19	b	20	b
21	d	22	b	23	c	24	c	25	d
26	c	27	c	28	c				

First law of thermodynamics and Hess law

1	d	2	b	3	a	4	b	5	c
6	a	7	c	8	a	9	b	10	d
11	b	12	c	13	c	14	c	15	b
16	b	17	c	18	b	19	c	20	a
21	c	22	c	23	a	24	a	25	b
26	a	27	a	28	a	29	bc	30	c
31	b	32	c	33	b	34	a	35	c
36	b	37	c	38	b	39	d	40	c

Heat of reaction

1	b	2	b	3	d	4	d	5	d
6	b	7	c	8	a	9	b	10	a
11	bcd	12	a	13	d	14	b	15	c
16	a	17	b	18	a	19	b	20	c
21	c	22	a	23	d	24	c	25	c
26	b	27	b	28	b	29	a	30	c
31	a	32	a	33	d	34	b	35	b
36	a	37	a	38	d	39	b	40	b
41	c	42	c	43	b	44	b	45	b
46	b	47	a	48	c	49	e	50	b
51	a	52	b	53	c	54	c	55	b
56	c	57	c	58	d	59	c	60	b
61	d	62	c	63	b	64	a	65	b
66	b	67	a	68	c	69	a	70	d
71	a	72	b	73	d	74	c	75	b
76	b	77	a	78	d	79	d	80	a
81	b	82	b	83	a	84	b	85	b
86	c	87	a	88	b	89	d	90	d
91	a	92	c	93	b	94	a	95	d
96	d	97	c	98	a	99	c	100	c
101	b	102	c	103	d	104	d	105	c
106	a	107	c	108	d	109	d	110	c
111	a	112	b	113	a	114	a	115	a
116	a	117	b	118	c	119	b	120	a
121	a	122	a	123	d	124	b	125	b
126	b	127	c	128	d	129	d	130	b
131	d	132	b	133	d	134	a	135	a
136	a	137	d	138	c	139	d	140	b
141	c	142	d	143	c	144	b	145	c
146	d	147	b	148	c	149	b	150	b
151	d	152	a	153	b	154	b	155	b
156	b	157	a	158	d	159	c	160	d
161	a	162	c	163	c	164	c	165	c
166	b	167	d	168	d	169	c	170	b
171	a	172	c						

Bond energy

1	b	2	a	3	a	4	d	5	c
6	c	7	b	8	d	9	b	10	b

Free energy and Work function

1	d	2	b	3	a	4	b	5	acd
6	c	7	c	8	c	9	a	10	a
11	d	12	a	13	d	14	d	15	b
16	d	17	b	18	c	19	d	20	b
21	b	22	d	23	b	24	a	25	d
26	d	27	a	28	d				

Critical Thinking Questions

1	b	2	c	3	d	4	b	5	b
6	d	7	d	8	b	9	a	10	a
11	a	12	c	13	c	14	a	15	a
16	a	17	a	18	d	19	b	20	a
21	b	22	c	23	c	24	b	25	b
26	b	27	e	28	d	29	a	30	c
31	a	32	c	33	d	34	ac	35	a
36	b	37	c	38	b	39	c	40	a
41	b	42	a	43	a	44	a	45	d
46	d	47	d	48	b	49	d	50	d
51	b	52	a						

Assertion & Reason

1	b	2	b	3	b	4	c	5	a
6	b	7	b	8	a	9	b	10	a
11	a	12	d	13	a	14	e	15	c
16	a	17	c	18	b	19	a	20	e
21	c	22	c						

AS Answers and Solutions

Basic concepts

- (b) Internal energy of an ideal gas is a function of temperature only.
- (d) In cyclic process, a system in a given state goes through a series of different processes, but in the end returns to its initial state.
- (c) $\Delta E = 0$ for isothermal reversible cycle.
- (c) In isolated system neither exchange of matter nor exchange of energy is possible with surroundings.

440 Thermodynamics and Thermochemistry

9. (c) It is the definition of calorific value.
11. (d) When a real gas is forced through a porous plug into a region of low pressure, it is found that due to expansion, the gas on the side of low pressure gets cooled.
The phenomenon of producing lowering of temperature when a gas is made to expand adiabatically from a region of high pressure into a region of low pressure is known as Joule-Thomson effect.
12. (d) In isothermal reversible process ideal gas has constant volume and so $\Delta E = 0$ and $\Delta H = \Delta E = 0$.
13. (a) The compressor has to run for longer time releasing more heat to the surroundings.
14. (c) This is based on Joule-Thomson effect.
15. (b) Enthalpy is an extensive property.
16. (b) $dV = 0$ for an isochoric process.
17. (d) Total energy of an isolated system is constant.
18. (a) For isochoric process $\Delta V = 0$ so $q_v = \Delta E$ i.e. heat given to a system under constant volume is used up in increasing ΔE .
19. (b) The less energy of a system and more is its stability.
20. (b) The functions whose value depends only on the state of a system are known as state functions.
21. (d) For adiabatic process $q = 0$.
22. (b) The intensive property is mass/volume.
23. (c) Volume is not an intensive property.
24. (c) An isolated system neither shows exchange of heat nor matter with surroundings.
25. (d) ΔQ is not a state function.
26. (c) For adiabatic process $\Delta Q = 0$.
27. (c) Work is not a state function as during a process its value depends on the path followed. The value of enthalpy, internal energy and entropy depends on the state and not on the path followed to get that state, hence these are state functions.
28. (c) Surface tension is an intensive property which do not depend upon the quantity of matter present in the system.
5. (c) $\Delta n_g = 1 - \frac{3}{2} = \frac{-1}{2}$, As Δn_g is negative, thus $\Delta H < \Delta E$.
6. (a) Bomb calorimeter is commonly used to find the heat of combustion of organic substances which consists of a sealed combustion chamber, called a bomb. If a process is run in a sealed container then no expansion or compression is allowed, so $w = 0$ and $\Delta U = q$.
 $\Delta U < 0, w = 0$
9. (b) $\Delta H = \Delta E + \Delta nRT$
Since $\Delta n = -2$
Then $\Delta H = \Delta E - 2RT$.
10. (d) If $\Delta n = -ve$ than $\Delta H < \Delta E$.
12. (c) Hess's law is an application of first law of thermodynamics.
13. (c) At constant P or T
 $\Delta H = \Delta U + \Delta nRT \Rightarrow \Delta n = n_p - n_R = 2 - 4 = -2$
 $\therefore \Delta H < \Delta U$.
16. (b) It is a combustion reaction, $\Delta H = -ve$.
17. (c) During isothermal expansion of ideal gas,
 $\Delta T = 0$
 $\Delta H = \Delta E + P\Delta V = \Delta E + nR\Delta T = 0 + 0 = 0$.
18. (b) $W = 2.303 nRT \log \frac{V_2}{V_1}$
 $= 2.303 \times 1 \times 8.314 \times 10^7 \times 298 \log \frac{20}{10}$
 $= 298 \times 10^7 \times 8.314 \times 2.303 \log 2$.
20. (a) The enthalpies of all elements in their standard state at $25^\circ C$ or $298K$ are zero.
21. (c) $\Delta E_v = E_p - E_R$.
22. (c) $\Delta E = q + w$.
23. (a) $\Delta E = 0$ for reversible isothermal process.
25. (b) Hess law includes initial reactants and final products.
26. (a) At constant T and P internal energy of ideal gas remains unaffected.
27. (a) ΔE increases with temperature.
28. (a) $\Delta H = \Delta E + W$ or $\Delta H = \Delta E + P\Delta V$.

First law of thermodynamics and Hess law

1. (d) First law of thermodynamics is also known as Law of conservation of mass and energy.
2. (b) Formation of CO_2 from CO is an exothermic reaction; heat is evolved from the system, i.e., energy is lowered. Thus, exothermic reactions occur spontaneously on account of decrease in enthalpy of system. Thus, $\Delta E > \Delta H$.
4. (b) $\Delta H = \Delta E + P\Delta V$.
29. (bc) Heat of neutralisation of a strong acid and strong base is equal to $-13.7 kcal$.
30. (c) $-W = +2.303 nRT \log \frac{p_1}{p_2}$
 $-W = 2.303 \times 1 \times 2 \times 300 \log \frac{10}{1} = 1381.8 cal$.
31. (b) Joule-Thomson expansion is isoenthalpic.
32. (c) Here $\Delta n = 0$ so, $\Delta E = \Delta H$.

33. (b) $q = \Delta E - W$ if $q = 0$ for adiabatic process, then $\Delta E = W$.
34. (a) For this reaction $\Delta n = 0$ than $\Delta E = \Delta H$.
35. (c) As the system is closed and insulated no heat enter or leave the system, i.e. $q = 0$; $\therefore \Delta E = Q + W = W$.
37. (c) $\Delta H - \Delta E = \Delta nRT$; $\Delta n = -3$
so, $\Delta H - \Delta E = -3RT$.
38. (b) According to Hess's law. The heat of reaction depends upon initial and final conditions of reactants.
39. (d) $\Delta H - \Delta E = \Delta nRT$
also, $2NH_{3(g)} \rightarrow N_{2(g)} + 3H_{2(g)}$
 $\Delta n = 2$.
40. (c) $\Delta n = 0$ for this reaction so, $\Delta E = \Delta H$.
41. (d) $W = 0$ is not true.
42. (a) $W = 2.303 nRT \log \frac{P_2}{P_1}$
 $= 2.303 \times 1 \times 2 \times 300 \log \frac{10}{2} = 965.84$
at constant temperature, $\Delta E = 0$.
 $\Delta E = q + w$; $q = -w = -965.84 \text{ cal}$.
43. (c) $q = 40 \text{ J}$
 $w = -8 \text{ J}$ (work done by the system)
 $\Delta E = q + w = 40 - 8 = 32 \text{ J}$.
44. (a) We know that $\Delta E = Q + W = 600 + (-300) = 300 \text{ J}$
 $W = 300$, because the work done by the system.
45. (c) Given that
 $P_1 = 10 \text{ atm}$, $P_2 = 1 \text{ atm}$, $T = 300 \text{ K}$, $n = 1$
 $R = 8.314 \text{ J/K/mol}$
Now, by using
 $W = 2.303 nRT \log_{10} \frac{P_2}{P_1}$
 $= 2.303 \times 1 \times 8.314 \times 300 \log_{10} \frac{1}{10}$
 $W = 5744.1 \text{ Joule}$
46. (c) We know that internal energy of a gas depends upon its pressure and temperature. Thus if a gas expands at constant temperature and pressure, then its internal energy remains same.
47. (b) $PCl_5(g) \rightarrow PCl_3(g) + Cl_2(g)$
For this reaction $\Delta n_g = 2 - 1 = 1$
 Δn_g is positive, i.e., there is an increase in the number of gaseous moles then $\Delta H > \Delta E$
48. (c) Enthalpy (H) is defined as the sum of internal energy $E + PV$, $H = E + PV$.
49. (d) $(CH_3)_2C = CH_2(g) + 6O_2(g) \rightarrow 4CO_2(g) + 4H_2O(l)$
 $\Delta n_g = 4 - 6 = -2$ (i.e., negative)
we know that $\Delta H = \Delta E + \Delta n_g RT$
 $= \Delta E - (\Delta n_g)RT$ ($\because \Delta n_g = -ve$)
 $\therefore \Delta H < \Delta E$
50. (b) Given number of moles = 1
Initial temperature = $27^\circ C = 300 \text{ K}$
Work done by the system = $3 \text{ KJ} = 3000 \text{ K}$
It will be (-) because work is done by the system.
Heat capacity at constant volume (C_V) = 20 J/k
We know that work done
 $W = -nC_V(T_2 - T_1)$; $3000 = -1 \times 20 (T_2 - 300)$
 $3000 = -20T_2 + 6000$
 $20T_2 = 3000$; $T_2 = \frac{3000}{20} = 150 \text{ K}$
51. (a) Internal energy of a system is a state function and extensive property and is independent of the path by which it is obtained.
52. (d) $N_2O_4(g) \rightarrow 2NO_2(g)$
For this reaction $\Delta n_g = 2 - 1 = 1$
 Δn_g is positive i.e., there is an increase in the number of gaseous moles then $\Delta H > \Delta E$
53. (b) $2C + O_2 \rightarrow 2CO$; $\Delta H = -220 \text{ KJ}$
This reaction does not represent complete combustion of carbon, hence heat of combustion of carbon will not be equal to 110 kJ . The negative sign of ΔH indicates that this reaction is exothermic. Also, despite being spontaneous reaction, it requires initiation.
54. (c) $W = -P\Delta V = -1 \times 10^5 (1 \times 10^{-2} - 1 \times 10^{-3})$
 $= -1 \times 10^5 \times 9 \times 10^{-3} = -900 \text{ J}$
55. (a) The exact value of internal energy is not known as it includes all type of energies of molecules constituting the given mass of matter such as translational, vibrational, rotational. The kinetic & potential energy of the nuclei & electron with in the individual molecules and the manner in which the molecule are linked together etc.
 $E = E_{\text{translational}} + E_{\text{rotational}} + E_{\text{vibrational}}$
Thus, we can say that internal energy is partly potential and partly kinetic.
56. (b) $N_2(g) + O_2(g) \rightarrow 2NO(g)$; $\Delta n_g = 2 - 2 = 0$
 $\Delta H = \Delta E + \Delta n_g RT$; $\Delta H = \Delta E + 0$; $\Delta H = \Delta E$
57. (d) $W = -p\Delta V$; $W = -3 \times (6 - 4)$
 $W = -6 \times 101.32$ ($\because 1 \text{ Latm} = 101.32 \text{ J}$)
 $W = -608 \text{ J}$
58. (d) $A(s) + 3B(g) \rightarrow 4C(s) + D(l)$
 $\Delta n_g = 0 - 3 = -3$; $\Delta H = \Delta E + \Delta n_g RT$ [$\because \Delta E = \Delta U$]
 $\therefore \Delta H = \Delta U + (-3)(RT)$
 $\Delta H = \Delta U - 3(RT)$; $\Delta H = \Delta U - 3RT$

442 Thermodynamics and Thermochemistry

2. (d) When $\Delta S = +ve$ the change is spontaneous.
3. (d) Heat is always flow from the higher to lower temperature.
4. (b) Mixing of non-reacting gases increases randomness and so increase entropy .
5. (b) Entropy of the system increases as the process occur irreversibly and isothermally in an isolated system.
6. (d) $\Delta S^\circ = 2S_{HCl}^\circ - (S_{H_2}^\circ + S_{Cl_2}^\circ)$
 $= 2 \times 186.7 - (130.6 + 223.0) = 19.8 \text{ JK}^{-1} \text{ mol}^{-1}$
9. (b) For adiabatic expansion $q = 0$ than according to following relation $\Delta S = \frac{q}{T}, \Delta S = 0$.
10. (c) It is the third law of thermodynamics.
11. (c) Entropy of gases is highest.
12. (c) $\Delta G = \Delta H - T\Delta S$
 for spontaneous process ΔG should be negative in option (3) $\Delta H = -ve$ and $\Delta S = +ve$ than
 $\Delta G = (-ve) - T(+ve) = -ve$
15. (c) For reverse reaction sign will be change.
16. (b) Solid \longrightarrow Gas, ΔS is maximum.
17. (d) $+ve \Delta H$ and $-ve \Delta S$ both oppose the reaction.
18. (a) $\Delta S_{vap} = 186.5 / 373 = 0.5 \text{ JK}^{-1} \text{ mol}^{-1}$.
20. (a) When $\Delta H = -ve$, $\Delta S = +ve$ and $\Delta G = -ve$ than reaction is spontaneous .
21. (d) For endothermic process ΔS increases.
22. (c) Calculation of change in entropy is done at constant temperature and pressure both.
23. (a) When the value of entropy is greater, then ability of work is maximum.
24. (d) At equilibrium, $\Delta G = 0$
 Hence $0 = \Delta H - T\Delta S$ or $\Delta H = T\Delta S$.
25. (c) $\Delta S_{vap} = \frac{\Delta H_{vap}}{T} = \frac{37.3 \text{ KJ mol}^{-1}}{373 \text{ K}}$
 $= 0.1 \text{ kJ mol}^{-1} \text{ K}^{-1} = 100 \text{ J mol}^{-1} \text{ K}^{-1}$.
26. (d) Solid \rightarrow Liquid, ΔS increases .
27. (c) $\Delta S = +ve$ than process is spontaneous.
28. (b) $\Delta S = +ve$ than the system is more disordered.
29. (c) Because solid \rightarrow solid, ΔS is same and ΔH is $-ve$.
30. (b) $\Delta S_{vap} = \frac{\Delta H_{vap}}{T} = \frac{386}{298} = 1.2 \text{ kJ}$.
31. (b) Processes (a) and (c) take place with the increase of no. of moles of gaseous species and hence the disorder or entropy increases.
 (b) on increasing pressure, disorder or randomness decreases and so also the entropy (d) is endothermic process and ΔS is positive.
32. (d) This is the statement of third law of thermodynamics.
33. (d) For isothermal expansion of ideal gas, $\Delta E = 0$.
34. (a) $\Delta S_{vap} = \frac{(900 \times 18)}{373} = 43.4 \text{ JK}^{-1} \text{ mol}^{-1}$.
35. (c) Entropy of the mixture increases due to increase in impurity.
37. (c) If $\Delta H = +ve$ and $\Delta S = -ve$ than the reaction is non-spontaneous i.e. can not occur.
38. (a) $T_2 = 150 + 273 = 423 \text{ K}$
 $T_1 = 25 + 273 = 298 \text{ K}$
 $Q = 500 \text{ K}$
 $\frac{W}{Q} = \frac{T_2 - T_1}{T_2}$; $W = 500 \left(\frac{423 - 298}{423} \right) = 147.7 \text{ J}$.
39. (c) According to III law of thermodynamics.
40. (c) When $\Delta H = +ve$ and $\Delta S = -ve$ reaction is non-spontaneous.
41. (c) $T_m = \frac{\Delta H_{fusion}}{\Delta S_{fusion}} = \frac{9.2}{0.008} = 1150 \text{ K}$.
42. (c) Here: Change in
 Volume (V) = $500 - 300 = 200 \text{ cc} = 0.2 \text{ litre}$,
 Pressure (P) = 0.6 atm and heat liberated (q) = 10 J
 Work done (W) = $P\Delta V = (0.2 \times 0.6) = 0.12 \text{ litre-atm}$
 But $1 \text{ litre-atm} = 101.3 \text{ J}$.
 hence $W = 0.12 \times 101.3 = 12.156 \text{ J}$. We also know that heat is liberated, therefore it would be negative. Thus change in $\Delta E = q + W = -10 + 12.16 = 2.16 \text{ J}$.
43. (c) Formation of CO_2 is,
 $C_{(s)} + O_{2(g)} \rightarrow CO_{2(g)}$
 $\Delta S^\circ = 213.5 - 5.690 - 205 = 2.81 \text{ JK}^{-1}$.
45. (b) $H_2O_{(g)} \rightleftharpoons H_2O_{(l)}$
 we know $\Delta G = \Delta H - T\Delta S$
 at equilibrium $\Delta G = 0$
 Therefore $\Delta H = T\Delta S$.
46. (a) We know that work done, $W = C_v(T_1 - T_2)$
 $3 \times 1000 = 20(300 - T_2)$; $\therefore 3000 = 6000 - 20 T_2$
 $\therefore T_2 = \frac{3000}{20} = 150 \text{ K}$.
48. (d) It does not violates the first law of thermodynamics but violates the II law of thermodynamics.
49. (d) $C_v = \frac{3}{2} RT$; $C_p = \frac{5}{2} RT$ for monoatomic gas
 $C_v = \frac{5}{2} RT$; $C_p = \frac{7}{2} RT$ for diatomic gas
 Thus for mixture of 1 mole each,
 $C_v = \frac{\frac{3}{2} RT + \frac{5}{2} RT}{2}$ and $C_p = \frac{\frac{5}{2} RT + \frac{7}{2} RT}{2}$

Therefore, $C_p / C_v = \frac{3RT}{2RT} = 1.5$.

51. (c) $\Delta S = \frac{q_{rev}}{T}$ \therefore unit of S is $JK^{-1}mol^{-1}$.

52. (b) $H_2O_{(l)} \rightleftharpoons H_2O_{(g)}$, $\Delta S = \frac{\Delta H_{vap}}{T}$,

$$\Delta H_{vap.} = 2.257 \text{ KJ / g}$$

$$\text{or } \Delta H_{vap} = 2.257 \times 18 \text{ kJ / mol.} = 40.7 \text{ KJ / mol}$$

$$\text{hence, } \Delta S = \frac{40.7}{373} = 0.109 \text{ kJ/mol/K.}$$

53. (a) Liquid \longrightarrow Vapour, entropy increases.

54. (c) $NaNO_3$ is a solid, which is converted to liquid ions.

55. (c) Heat capacity of water per gram = $\frac{75}{18} = 4.17 \text{ J}$

$$Q = mst ; 1000 = 100 \times 4.17 \times t$$

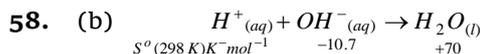
$$t = \frac{1000}{100 \times 4.17} = 2.4 \text{ K.}$$

56. (c) As the work is done on system, it will be positive i.e. $W = +462 \text{ joule}$, $E = -128 \text{ joule}$ (heat is evolving)

From the 1st law of thermodynamics

$$\Delta E = q + w = (-128) + (+462) = +334 \text{ Joules.}$$

57. (a) Gases show highest entropy.



$$\Delta S^\circ(298 \text{ K}) = \Delta S_p - \Delta S_R = 70 - (-10.7 + 0)$$

$$80.7 \text{ JK}^{-1} \text{ mol}^{-1}$$

59. (d) The entropy change = $\frac{\text{heat of vaporisation}}{\text{temperature}}$

Here, heat of vaporisation = 540 cal / gm

$$= 540 \times 18 \text{ cal mol}^{-1}$$

Temperature of water = $100 + 273 = 373 \text{ K}$

\therefore entropy change

$$= \frac{540 \times 18}{373} = 26.06 \text{ cal mol}^{-1} \text{ K}^{-1}$$

60. (b) Given that, $T_1 = 500 \text{ K}$, $T_2 = 300 \text{ K}$

$$\text{By using, } \eta = \frac{T_1 - T_2}{T_1} = \frac{500 - 300}{500} = \frac{200}{500} = 0.4 .$$

61. (b) It is molar heat capacity.

62. (d) Entropy is the measure of randomness in the molecules. Randomness is maximum in case of gases. Hence, entropy is maximum for water vapours.

63. (a) Standard entropy of formation of $CO_2(g)$ = standard entropy of $CO_2(g)$ - [Standard entropy of $C(s)$ - standard entropy of $O_2(g)$]
 $= 213.5 - [5.740 + 205] = 2.76 \text{ J / K.}$

64. (c) Third law of thermodynamics help in calculating entropy of different temperatures.

65. (a) In case of gas randomness is maximum therefore entropy is maximum in case of steam.

66. (c) $\Delta S_{system} + \Delta S_{surroundings} > 0$ (for spontaneity)

$$\text{Because of } \Delta S = R \ln \frac{V_2}{V_1}$$

Here the volume of gas increase from V_1 to V_2 at constant temperature T .

The total increase in entropy of the system and its surrounding during the spontaneous process of expansion considered above is,

thus $R \ln \left(\frac{V_2}{V_1} \right)$ since $V_2 > V_1$ it is obvious that

the spontaneous (irreversible) isothermal expansion of a gas is accompanied by an increase in the entropy of the system and its surrounding considered together.

$$\Delta S_{system} + \Delta S_{surroundings} > 0 .$$

67. (b) $\Delta G = \Delta H - T\Delta S$

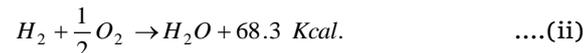
at constant temperature and pressure $\Delta G = 0$

$$0 = \Delta T - T\Delta S \text{ so } \Delta H = T\Delta S$$

68. (b) $dS = \frac{dQ_{rev.}}{T}$; $T = \frac{30 \times 10^3}{75}$; $T = 400 \text{ K}$

Heat of reaction

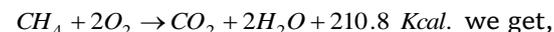
1. (b) $C + O_2 \rightarrow CO_2 + 94.2 \text{ Kcal.}$ (i)



On multiplication of eq. (ii) by 2 and then adding in eq. (i)



On subtracting eq. (iii) by following eq.



2. (b) $\Delta S = 16 \text{ J mole}^{-1} \text{ K}^{-1}$

$$T_{b.p.} = \frac{\Delta H_{vapour}}{\Delta S_{vapour}} = \frac{6 \times 1000}{16} = 375 \text{ K}$$

4. (d) Heat of neutralisation between strong acid and a strong base is about -13.7 Kcal.

6. (b) Effect of temperature in heat of reaction is given by Kirchoff's equation.

7. (c) Heat of neutralisation between strong acid and a strong base is about -13.7 Kcal.

8. (a) For exothermic reactions $H_p < H_R$.

For endothermic reactions $H_p > H_R$.

9. (b) 78g of benzene on combustion produces heat = -3264.6 kJ

$$\therefore 39 \text{ g will produce} = \frac{-3264.6}{2} = -1632.3 \text{ kJ.}$$

10. (a) eq. (i) + eq. (ii) gives the required result.

444 Thermodynamics and Thermochemistry

11. (bcd) (b,c,d) are endothermic reactions because they proceed by the absorption of heat.
12. (a) Change of liquid to vapour takes energy in the form of heat so it is endothermic reaction.
13. (d) In exothermic reactions heat is evolved.
15. (c) $\Delta H = -ve$ for exothermic reaction.
 $\Delta H = +ve$ for endothermic reaction
 Enthalpy of fusion is $+ve$.
16. (a) Heat of neutralisation will be less than -57.33 kJ/mole because some amount of this energy will be required for the dissociation of weak base (MgO)
18. (a) $H_2 + O_2 \rightarrow H_2O_2 \quad \Delta H_f^\circ = -188 \text{ kJ/mole} \quad \dots(i)$
 $H_2 + \frac{1}{2}O_2 \rightarrow H_2O \quad \Delta H_f^\circ = -286 \text{ kJ/mole} \quad \dots(ii)$
 eq. (i) - eq. (ii) $\times 2$ gives the required result.
19. (b) Graphite \longrightarrow diamond $\Delta H_t = (x-y) \text{ kJ mol}^{-1}$.
21. (c) Heats of combustions are always exothermic except oxidation of N as,
 $N_2 + \frac{1}{2}O_2 \rightarrow N_2O \quad \Delta H = +ve$
 $N_2 + O_2 \rightarrow 2NO \quad \Delta H = +ve$
22. (a) For exothermic reactions $H_p < H_R$.
 For endothermic reactions $H_p > H_R$.
23. (d) Aim: $2C + H_{2(g)} \rightarrow C_2H_{2(g)}$.
 eq. (ii) + eq. (iii) \rightarrow eq. (iv) - eq. (i)
 find the required result.
24. (c) Enthalpy of formation of HCl .
25. (c) Heat of neutralisation between strong acid and a strong base is about -13.7 Kcal .
26. (b) For endothermic reaction, $\Delta H = +ve$.
27. (b) Heat of neutralisation is less than 56.1 Kcal when a strong base and a weak acid reacts.
28. (b) Aim: $CO + \frac{1}{2}O_2 \rightarrow CO_2$

$$\Delta H = \Delta H_f^\circ(CO_2) - \left[\Delta H_f^\circ(CO) + \frac{1}{2} \Delta H_f^\circ(O_2) \right]$$

$$= -94.0 - (-26.4) = -67.6 \text{ kcal}.$$
29. (a)
 (i) $H_2 + \frac{1}{2}O_2 \rightarrow H_2O, \Delta H = -241 \text{ kJ}$
 (ii) $C_6H_{10} + \frac{17}{2}O_2 \rightarrow 6CO_2 + 5H_2O, \Delta H = -3800 \text{ kJ}$
 (iii) $C_6H_{12} + 9O_2 \rightarrow 6CO_2 + 6H_2O, \Delta H = -3920 \text{ kJ}$
 $C_6H_{10} + H_2 \rightarrow C_6H_{12}$
 Eq. (i) + Eq. (ii) - Eq. (iii)
 $\Delta H = -241 - 3800 - (-3920)$
 $= -4041 + 3920 = -121 \text{ kJ}$
30. (c) NH_4OH is a weak base. Heat of neutralisation $< 13.7 \text{ kcal}$.
31. (a) CH_4 is the best fuel because its calorific value $= \frac{-212.8}{16} = -13.3 \text{ kcal/g}$ is higher among the other gases.
32. (a) Find ΔH for $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O_{(g)}$.
34. (b) In neutralisation of a strong acid and base only H^+ and OH^- ions react.
35. (b) When both acid and base are strong than heat of neutralisation is 57.1 kJ mol^{-1} .
36. (a) $C_2H_6 + \frac{7}{2}O_2 \rightarrow 2CO_2 + 3H_2O$ substitute the values.
37. (a) Decomposition is an endothermic process.
38. (d) ΔH for this reaction is Heat of reaction.
40. (b) One mole of H_2O is formed from its initial components.
41. (c) One mole of a substance is completely oxidized in air.
42. (c) S (rhombic) $+ O_2 \rightarrow SO_2$, $\Delta H = 70960 \text{ cal}$ (i)
 S (monoclinic) $+ O_2 \rightarrow SO_2$, $\Delta H = 71030 \text{ cal}$... (ii)
 Aim: S (rhombic) $\rightarrow S$ (monoclinic)
 eq. (i) - eq. (ii) gives the required result.
43. (b) When $H_2O_{(l)}$ is converted to form $H_2O_{(g)}$, heat is absorbed hence $\Delta H_1 > \Delta H_2$.
44. (b) Out of given substances, kerosene oil has maximum calorific value.
45. (b) $C_{(s)} + O_{2(g)} \rightarrow CO_{2(g)}$, $\Delta n_g = 1 - 1 = 0$
 $\therefore \Delta E = \Delta H = -94 \text{ kcal}$
 $\Delta E = E_{CO_2} - (E_C + E_{O_2}) = E_{CO_2} - (0 + 0)$
 or $E_{CO_2} = \Delta E = -94 \text{ kcal}$.
46. (b) 0.2 mole will neutralize 0.2 mole of HNO_3 , heat evolved $= 57 \times 0.2 = 11.4 \text{ kJ}$.
47. (a) Suppose heat evolved in Ist case is Q_1 and that in the IInd case it is Q_2 . Then $Q_2 = \frac{1}{2}Q_1$.
 But $Q_1 = 1000 T_1$ and $Q_2 = 500 T_2$
 $\therefore 500 T_2 = \frac{1}{2} \times 1000 T_1$ i.e. $T_2 = T_1$.
48. (c) Enthalpy of a compound = Heat of reaction of that compound.
49. (e) It is the definition of heat of neutralization.
50. (b) $\Delta H = -ve$ for exothermic compound.
51. (a) If acid or base or both are strong, heat of neutralization $= 13.7 \text{ kcal}$.
52. (b) Both NH_4OH and CH_3COOH are weak.
53. (c) $57.1 \times 0.25 = 14.3 \text{ kJ mol}^{-1}$.

54. (c) Heat of formation is for 1 mole.
Hence $\Delta H_f^0(HI) = -12.40/2 = -6.20 \text{ kcal}$.
55. (b) Chemical dissociations are reversible and endothermic.
57. (c) Given,
(i) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$, $\Delta H = +20 \text{ kcal}$
(ii) $C + O_2 \rightarrow CO_2$, $\Delta H = -40 \text{ kcal}$.
(iii) $H_2 + \frac{1}{2}O_2 \rightarrow H_2O$, $\Delta H = -10 \text{ kcal}$.
Aim: $C + 2H_2 \rightarrow CH_4$
(ii) + 2 × (iii) - (i) gives.
 $\Delta H = -40 + 2(-10) - (+20) = -80 \text{ kcal}$.
58. (d) For exothermic reaction heat is evolved.
59. (c) $H_2O(l) \xrightarrow{\text{electrolysis}} H_{2(g)} + \frac{1}{2}O_{2(g)}$.
61. (d) All other are combustion phenomena.
62. (c) $\Delta H_{\text{combustion}} = -ve$, so exothermic process.
63. (b) Aim: $K_{(s)} + \frac{1}{2}O_{2(g)} + \frac{1}{2}H_{2(g)} \rightarrow KOH_{(s)}$
eq. (ii) + eq. (i) - eq. (iii) gives
 $\Delta H = -48 + (-68.39) - (-14) = -68.39 - 48 + 14$.
64. (a) By definition of heat of formation.
68. (c) $C_D \rightarrow C_G$, $\Delta H = -453.5 \text{ cal}$.
i.e. energy of C_G is less and thus more stable.
69. (a) $\Delta H_{\text{combustion}}$ is always negative.
70. (d) $\Delta H_{\text{reaction}} = \Delta H_f^0(N_2O_4) - 2\Delta H_f^0(NO_2)$
 $= 2 - 2(8) = -14 \text{ kcal}$.
71. (a) $\Delta H - \Delta E = \Delta nRT = 3 \times 8.314 \times 298$
 $= -7432 \text{ J} = -7.43 \text{ kJ}$.
72. (b) Aim: $C_{(\text{graphite})} \rightarrow C_{(\text{diamond})}$; eq. (i) - (ii) gives the result.
73. (d) In the formation of a compound, more the heat absorbed, less stable is the compound.
74. (c) $C_{(\text{graphite})} + 2H_{2(g)} = CH_{4(g)}$.
75. (b) Heat of formation is the formation of one mole of the substance from its elements.
76. (b) $\Delta H_{(\text{reaction})} = 2\Delta H_f^0(MgO) - \Delta H_f^0(SiO_2)$
 $= 2(-34.7) - (-48.4) = -21 \text{ kJ}$.
77. (a) eq. (i) + eq. (ii) + eq. (iii) gives
 $X = 131 - 282 - 242 = -393 \text{ kJ}$.
78. (d) It is the definition of heat of transition.
79. (d) $C_{(\text{graphite})} \rightarrow C_{(\text{diamond})}$, $\Delta H = 1.9 \text{ kJ}$
 $C_{(\text{graphite})} + O_2 \rightarrow CO_2$, $\Delta H = -\Delta H_1$
 $C_{(\text{diamond})} + O_2 \rightarrow CO_2$, $\Delta H = -\Delta H_2$
 $(-\Delta H_1) - (-\Delta H_2) = 1.9 \text{ kJ}$ or $\Delta H_2 = \Delta H_1 + 1.9$
For combustion of 6g, $\Delta H_2 > \Delta H_1$ by
 $1.9/2 = 0.95 \text{ kJ}$.
80. (a) $H_{\text{product}} \times H_{\text{reactant}}$ for exothermic reaction.
81. (b) Due to high electron affinity of Cl the highest energy is evolved.
82. (b) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$
Molecular weight of $CH_4 = 12 + 4 = 16$
 \therefore On the combustion of 2.0gm of methane = 25.0 kcal
 \therefore On the combustion of 16.0gm = $\frac{25 \times 16}{2} = 200 \text{ kcal}$.
83. (a) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$.
84. (b) $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
 $\Delta H_{\text{reaction}} = [2 \times \Delta H_f^0(CO_2) + 2 \times \Delta H_f^0(H_2O)]$
 $- [\Delta H_f^0(C_2H_4) + 3 \times \Delta H_f^0(O_2)]$
 $= [2(-394) + 2(-286)] - [52 + 0] = -1412 \text{ kJ}$.
85. (b) $C + O_2 \rightarrow CO_2 + 394 \text{ kJ}$.
 $2C + 2O_2 \rightarrow 2CO_2 + 788 \text{ kJ}$ (i)
 $2CO + O_2 \rightarrow 2CO_2 + 569 \text{ kJ}$ (ii)
 $2CO_2 \rightarrow 2CO + O_2 - 569 \text{ kJ}$ (iii)
eq. (i) + eq. (iii) = $-109.5 \text{ kJ mol}^{-1}$.
86. (c) $\Delta H_f = \frac{44}{2} \text{ kcal} = 22 \text{ kcal}$.
87. (a) $CaO_{(s)} + H_2O_{(l)} \rightarrow Ca(OH)_2$ is exothermic.
88. (b) $\Delta H = -ve$ in exothermic reaction.
89. (d) According to the definition of heat of formation.
90. (d) Multiplying eq. (ii) by (iii) and eq. (iii) by (vi), and then add
 $6C + 3H_2 + \frac{15}{2}O_2 \rightarrow 6CO_2 + 6H_2O$; $\Delta H = 3218.7$
Subtract eq. (i) from the above equation and find the required result.
91. (a) HF is more stable than HCl.
92. (c) $\Delta H(H_P - H_R) = q_p$.
93. (b) $\Delta H = +ve$ for endothermic reaction.
94. (a) $\Delta H = \frac{13.95 \times 44}{2.2016} = 278.7 \text{ kcal}$.
95. (d) eq. (i) + eq. (ii) gives the required result.
96. (d) Standard molar heat enthalpy (H°) of a compound is equal to its standard heat of formation from most stable states of initial components.
97. (c) In the complete combustion of butanol $\Delta H > \Delta E$.
98. (a) X - Y and find the required result.
99. (c) $S + O_2 \rightarrow SO_2$ $\Delta H = -298.2 \text{ kJ}$ (i)
 $SO_2 + \frac{1}{2}O_2 \rightarrow SO_3$ $\Delta H = -98.2 \text{ kJ}$ (ii)
eq. (i) - (ii) and find the required result.
100. (c) $C + 2S \rightarrow CS_2$ $\Delta H = ?$
 $C + O_2 \rightarrow CO_2$ $\Delta H = -393.3 \text{ kJ}$ (i)
 $S + O_2 \rightarrow SO_2$ $\Delta H = -293.72 \text{ kJ}$ (ii)

446 Thermodynamics and Thermochemistry

- $CS_2 + 3O_2 \rightarrow CO_2 + 2SO_2 \quad \Delta H = -1108.76 \text{ kJ} \dots(\text{iii})$
 $\text{eq. (ii)} \times 2 + \text{eq. (i)} - (\text{iv}) \text{ eq. (iii)} - \text{eq. (iv)}$
 and find required result .
- 101.** (b) Fermentation is exothermic reaction.
- 102.** (c) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$.
- 103.** (d) Use $\Delta H = \Delta E + \Delta n RT$
 $\Delta H = 19 + 2 \times 2 \times 10^{-3} \times 300 = 20.2 \text{ kcal} ; \Delta n = 2$.
- 104.** (d) $-20.6 - 8.8 \text{ KJ mol}^{-1} = -29.4 \text{ kJ}$.
- 106.** (a) Subtract equation (ii) from (i).
- 107.** (c) $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O, \Delta H = -890.4 \text{ kJ} \dots(\text{i})$
 $C + O_2 \rightarrow CO_2, \Delta H = -395.5 \text{ kJ} \dots(\text{ii})$
 $C + O_2 \rightarrow CO_2, \Delta H = -395.5 \text{ kJ} \dots(\text{ii})$
 $H_2 + \frac{1}{2} O_2 \rightarrow H_2O, \Delta H = -285.8 \text{ kJ} \dots(\text{iii})$
 from (i), (ii), (iii).
 $\Delta H_{\text{comb.}}(CH_4)$
 $= \Delta H_f(CO_2) + 2\Delta H_f(H_2O) - \Delta H_f(CH_4) - 2\Delta H_f(O_2)$
 $= -890.4 = -395.5 + 2(-285.5) - \Delta H_f(CH_4) - 2 \times 0$
 $\Delta H_f(CH_4) = -76.7 \text{ kJ mol}^{-1}$.
- 108.** (d) As methanoic acid is weak acid, heat of neutralization $< x$.
- 109.** (d) Neutralization of a strong acid by a strong base is always same.
- 110.** (c) $S + O_2 \rightarrow SO_2, \Delta H_f = -4.6 \text{ kJ}$
 $\therefore 0.5 \text{ gm. of sulphur on burning produce } 1 \text{ gm of } SO_2$
 $\therefore 32 \text{ gm of sulphur on burning produce } 64 \text{ gm. of } SO_2$
 $\therefore \Delta H = (-4.6 \text{ kJ}) \times 64 = -294.4 \text{ kJ}$.
- 111.** (a) $3O_2 = 2O_3$ - energy is given out.
- 112.** (b) $\Delta H \text{ per } 1.6 \text{ g} = \frac{72 \times 1.6}{180} = 0.64 \text{ kcal}$.
- 113.** (a) $C + O_2 \rightarrow CO_2, \Delta H = -394 \text{ kJ} \dots(\text{i})$
 $2H_2 + O_2 \rightarrow 2H_2O, \Delta H = -568 \text{ kJ} \dots(\text{ii})$
 $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O, \Delta H = -892 \text{ kJ} \dots(\text{iii})$
 (i) + (ii) - (iii) and find the required result.
- 115.** (a) No doubt heat evolved during neutralisation of 250 cm^3 of each acid and base is five times the heat evolved during neutralisation of 50 cm^3 of each acid and base but the quantity of solution taking heat is also five times thus same temperature rise is noticed.
- 116.** (a) $H_2 + \frac{1}{2} O_2 \rightarrow H_2O_{(l)}; \Delta H = -285.8 \text{ KJ}$
 $H_2O_{(l)} \rightarrow H_2O_{(g)}; \Delta H = 44 \text{ KJ}$
 $\therefore H_2 + \frac{1}{2} O_2 \rightarrow H_2O_{(g)}; \Delta H^\circ = -241.8 \text{ KJ}$
- 117.** (b) Given: Weight of iron burnt = 4 g; Heat liberated = 29.28 KJ and atomic weight of iron (Fe) = 56 . We know that in ferric oxide (Fe_2O_3), 2 moles of iron or $2 \times 56 = 112 \text{ gram}$ of iron are burnt. We also know that when 4 grams of iron are burnt, then heat liberated = 29.28 kJ, therefore when 112 grams of the iron are burnt, then heat liberated = $\frac{29.28 \times 112}{4} = -819.8 \text{ kJ}$ (Minus sign due to liberation of heat).
- 118.** (c) $H^+ + OH^- \rightarrow H_2O, \Delta H_{\text{neutralization}} = 13.7 \text{ kcal}$.
- 119.** (b) $\Delta H / \text{mole of } FeS = \frac{3.77 \times 56}{2.1} = 100.5$.
- 120.** (a) Heat of formation = $\frac{194}{2} = 97 \text{ kJ}$.
- 121.** (a) $\Delta H = \Delta H_{\text{ioniz}} + \Delta H_{\text{neu}} = -50.6 = \Delta H_{\text{ioniz}} + (-55.9)$
 $\Delta H_{\text{ioniz}} = +5.3 \text{ kJ / mol}$.
- 122.** (a) Strong acid (HNO_3) and strong base ($LiOH$) .
- 123.** (d) $\Delta H = \Delta E + \Delta n RT$ or $\Delta E = \Delta H - \Delta n RT$
 $\therefore \Delta E = +7.3 - \frac{1}{2} \times 0.002 \times 298 = 7.3 - 0.298 = 7 \text{ kcal}$.
- 124.** (b) Heat evolved during burning of 39 g
 $C_6H_6 = \frac{781.0 \times 39}{78} = 390.5 \text{ kcal mol}^{-1}$
- 125.** (b) By (i) + (ii) ; $Na + \frac{1}{2} Cl_2 \rightarrow NaCl, \Delta H = -196 \text{ kcal}$.
- 126.** (b) $C + O_2 \rightarrow CO_{2(g)}, \Delta H_f^\circ = -393.5 \text{ kJ mol}^{-1} \dots(\text{i})$
 $C + \frac{1}{2} O_2 \rightarrow CO_{(g)}, \Delta H_f^\circ = -110.5 \text{ kJ mol}^{-1} \dots(\text{ii})$
 $H_2 + \frac{1}{2} O_2 \rightarrow H_2O_{(g)}, \Delta H_f^\circ = -241.8 \text{ kJ mol}^{-1} \dots(\text{iii})$
 By (ii) + (iii) - (i)
 $CO_{2(g)} + H_{2(g)} \rightarrow CO_{(g)} + H_2O_{(g)}, \Delta H = +41.2$.
- 127.** (c) More +ve is ΔH_s more is heat of solution.
- 128.** (d) CS_2 is formed from its initial components carbon and sulphur so, ΔH is heat of formation of CS_2 .
- 129.** (d) $\Delta H = 18500 = \Delta E + \Delta n RT$
 or $18500 = \Delta E + (-1) \times 2 \times 298 = 19096 \text{ cal}$.
- 130.** (b) Heat evolved during combustion of 0.39 g
 $C_6H_6 = \frac{3250 \times 0.39}{78} = 16.25 \text{ kJ}$.
- 131.** (d) By (i) - (ii) : $C + \frac{1}{2} O_2 \rightarrow CO; \Delta H = -229.2$.
- 132.** (b) Strong acid (HCl) and strong base ($NaOH$) shows $\Delta H_{\text{neutralization}} = -57.3 \text{ kJ}$.
- 133.** (d) By (i) - (ii) and find required result.

134. (a) $CO_{(g)} + \frac{1}{2}O_{2(g)} \rightarrow CO_{(g)}, \Delta n = 1 - 1 + \frac{1}{2} = -\frac{1}{2}$
 $\Delta H = \Delta E + \Delta nRT$
 $\Delta H = -283.3 - \frac{1}{2} \times \frac{8.314}{1000} \times 290 = -284.5 kJ$.
135. (a) $H_2 + \frac{1}{2}O_2 \rightarrow H_2O \quad \Delta H = -188 kJ/mole \quad \dots\dots(i)$
 $H_2 + O_2 \rightarrow H_2O_2; \Delta H = -286 kJ/mole \quad \dots\dots(ii)$
 By $2 \times (i)$ and (ii)
 $2H_2 + O_2 \rightarrow 2H_2O; \Delta H = -376 kJ/mole \quad \dots\dots(iii)$
 $2H_2 + 2O_2 \rightarrow 2H_2O_2 \quad \Delta H = -572 kJ/mole \quad \dots\dots(iv)$
 By $(iii) - (iv)$
 $2H_2O_2 \rightarrow 2H_2O + O_2 \quad \Delta H = +196 kJ$.
136. (a) ΔH for $C_2H_6 = -341.1 Kcal$
 its calorific value $= \frac{-341.1}{30} = -11.37 kcal/g$.
 ΔH for $C_2H_2 = -310.0 kcal$
 its Calorific value $= \frac{-310.0}{26} = -11.92$,
 hence C_2H_2 is a better fuel.
138. (c) For the decomposition of 9gm of water heat required = 142.5 kJ
 we know $H_2O = 2 + 16 = 18$
 Therefore heat required for decomposition of 18gm water $= \frac{18}{9} \times 142.5 = 285 KJ$
 Then, enthalpy of formation of water is reverse of heat required = - 285 kJ.
139. (d) $C_6H_{6(g)} + \frac{15}{2}O_{2(g)} \rightarrow 6CO_{2(g)} + 3H_2O_{(g)}$
 $\Delta n = 6 + 3 - 1 - \frac{15}{2} = +\frac{1}{2}$.
140. (b) By $2 \times (i) + (ii) - (iii)$
 ΔH of methane = 20.0 kcal.
141. (c) Heat of neutralization of strong acid and weak base is less than $13.7 kcal mol^{-1}$.
142. (d) $\Delta E = 0$ for a cyclic process.
143. (c) $22.0 \times 2 = 44 kcal$.
144. (b) 0.3 mole OH^- , neutralize 0.3 mole of HNO_3
 Evolved heat, = $57.1 \times 0.3 = 17.13 kJ$
146. (d) Compounds with high heat of formation are less stable because energy rich state leads to instability.
147. (b) CH_4 required $= \frac{445.15 \times 16}{890.3} = 8 gm$.
148. (c) $C_{(s)} + 2H_{2(g)} \rightarrow CH_{4(g)} \quad \dots\dots(i)$
 $C_{(s)} + O_{2(g)} \rightarrow CO_{2(g)} \quad \Delta H = -94 kcal mol^{-1} \quad \dots\dots(ii)$
 $H_{2(g)} + \frac{1}{2}O_{2(g)} \rightarrow H_2O_{(l)} \quad \Delta H = -68 kcal mol^{-1} \quad \dots\dots(iii)$
 $CH_4 + 3\frac{1}{2}O_2 \rightarrow CO_2 + 2H_2O$
 $\Delta H = -213 kcal mol^{-1} \quad \dots\dots(iv)$
 to obtain equation (i) operate $-(ii) + 2 \times (iii) - (iv)$.
149. (b) $\Delta H = +ve$ for endothermic reactions.
150. (b) It pertains to neutralization of strong acid weak base.
151. (d) Strong base (KOH) and strong acid (HCl)
153. (b) This reaction absorbed heat, so it is endothermic reaction.
154. (b) $C + O_2 \rightarrow CO_2, \Delta H = -94.3$
 This is also heat of formation of CO_2
 $C + \frac{1}{2}O_2 \rightarrow CO, \Delta H = -26.0$.
155. (b) $X = \frac{1}{2}Y$.
157. (a) Molecular weight of $NH_4NO_3 = 80$
 Heat evolved = 1.23×6.12
 \therefore Molar heat capacity = $1.23 \times 6.12 \times C$.
158. (d) Both X_e and F_2 are gaseous elements at $25^\circ C$ and in their standard states and form $XeF_{4(g)}$
 hence $\Delta H_{(f)}^\circ = \Delta H_{react}^\circ$.
159. (c) According to Hess law, enthalpy change for a reaction does not depend on the nature of intermediate reaction steps.
160. (d) By $(ii) - (i), C_{gr} \rightarrow C_{dia}, \Delta H = +1.9$.
162. (c) Heat of neutralization of strong acid and strong base is equal to the $-57.32 KJ mol^{-1}$
163. (c) $C + O_2 \rightarrow CO_2 \quad \Delta H = -393.5 KJ/mol$
 \therefore 44 gm of CO_2 form by which heat released = $-393.5 kJ$
 \therefore 1 gm of CO_2 form by which heat released
 $= -\frac{393.5}{44}$
 \therefore 35.2 gm (given) of CO_2 form by which heat released
 $= -\frac{393.5}{44} \times 35.2 = -315 kJ$
164. (c) $C_{(graphite)} + 2H_{2(g)} \rightarrow CH_{4(g)}$
165. (c) Heat of formation of a substance is the heat exchanged when one mole of that substance is formed by its constituent elements.

448 Thermodynamics and Thermochemistry

\therefore Heat evolved when 1 mole (44 g) CO_2 is formed = 393 kJ

\therefore Heat evolved when 0.156 Kg (156 g) is formed = $\frac{393 \times 156}{44}$

$\therefore \Delta H$ for the process = 1572 kJ = -1572.0 kJ

166. (b) It is heat of neutralisation (13.7 Kcal) for strong acid and strong base.

167. (d) $C(s) + O_2(g) \rightarrow CO_2(g)$

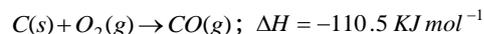
$$\Delta H = -393.5 \text{ KJ mol}^{-1} \quad \dots\dots(i)$$



$$\Delta H = -283 \text{ KJ mol}^{-1} \quad \dots\dots(ii)$$

On subtracting equation (ii) from equation (i)

We get

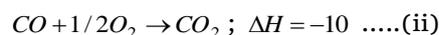
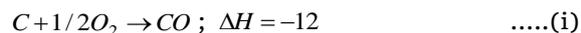


The enthalpy of formation of carbon monoxide per mole = -110.5 KJ mol⁻¹

168. (d) 1 mole (i.e.,) 16 gm of methane on combustion liberate 890 kJ

$$\therefore 3.2 \text{ gm will liberate} = \frac{890 \times 3.2}{16} = 178 \text{ KJ}$$

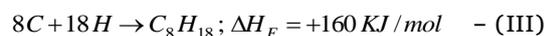
169. (c) $C + O_2 \rightarrow CO_2; \Delta H = q$



adding equation (i) and (ii) we can get

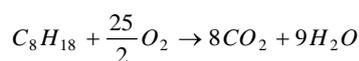
$$\Delta H = -12 + (-10) = -22$$

170. (b) $C + O_2 \rightarrow CO_2; \Delta H_F = -490 \text{ KJ/mol}$ - (I)



applying

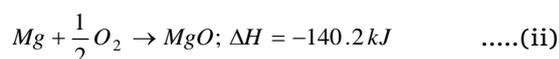
$$(I) \times 8 + (II) \times 9 + (III)$$



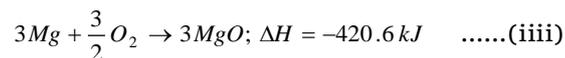
$$\Delta H^\circ = -3920 - 2160 - 160 = 6240 \text{ KJ/mol}$$

$$\Delta H^\circ = \text{for 6 moles of octane} = 6240 \times 6 = 37440 \text{ KJ/mol} = -37.4 \text{ KJ}$$

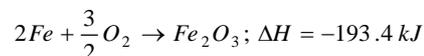
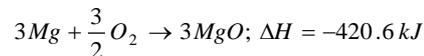
171. (d) $2Fe + \frac{3}{2} O_2 \rightarrow Fe_2O_3; \Delta H = -193.4 \text{ kJ}$ (i)



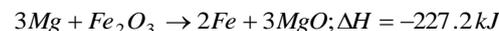
On multiplying eq. (ii) by 3



Resulting equation can be obtained by subtracting eq. (i) from (iii)



Subtraction :



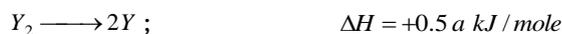
172. (c) HCl and KOH both are strong

Bond energy

1. (b) $XY \longrightarrow X_{(g)} + Y_{(g)}; \Delta H = +a \text{ kJ/mole}$ (i)

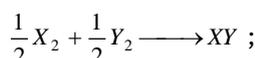


(ii)



.....(iii)

$$\frac{1}{2} \times (ii) + \frac{1}{2} \times (iii) - (i), \text{ gives}$$



$$\Delta H = \left(+\frac{a}{2} + \frac{0.5}{2} a - a \right) \text{ kJ/mole}$$

$$+\frac{a}{2} + \frac{0.5a}{2} - a = -200$$

$$a = 800$$

2. (a) 4 g $H_2 = 2$ moles. Bond energy for 1 mole of $H_2 = 208 / 2 = 104 \text{ kcal}$.

4. (d) By definition of Hess's law.

5. (c) Aim: $\frac{1}{2} H_2 + \frac{1}{2} Cl_2 \rightarrow HCl$

$$\Delta H = \sum B.E.(\text{Products}) - \sum B.E.(\text{Reactants})$$

$$= B.E.(HCl) - \left[\frac{1}{2} B.E.(H_2) + \frac{1}{2} B.E.(Cl_2) \right]$$

$$= -103 - \left[\frac{1}{2}(-104) + \frac{1}{2}(-58) \right]$$

$$= -103 - (-52 - 29) = -22 \text{ kcal}.$$

6. (c) First ionization potential of Li = 5.4 eV
Electron affinity of Cl = 3.61 eV

We have $\Delta H = I.P. - E.A.$

$$= 5.4 - 3.61 = 1.80 \text{ eV} = 1.80 \times 1.6 \times 10^{-22} \text{ kJ}$$

$$\Delta H = 2.86 \times 10^{-22} \text{ kJ},$$

for 1 mol Avogadro's number = 6.02×10^{23} .

$$\Delta H = 2.86 \times 10^{-22} \times 6.02 \times 10^{23} = 170 \text{ kJ/mole}.$$

7. (b) $\frac{-166}{4} = -41.5 \text{ kJ/mole}.$

8. (d) $\frac{1}{2}H_2 + \frac{1}{2}Cl_2 \rightarrow HCl, \Delta H = -90 \text{ KJ}$

$$\therefore \Delta H = \frac{1}{2}E_{H-H} + \frac{1}{2}E_{Cl-Cl}$$

$$\text{or } -90 = \frac{1}{2} \times 430 + \frac{1}{2} \times 240 - E_{HCl}$$

$$\therefore E_{H-Cl} = 425 \text{ kJ mol}^{-1}.$$

9. (b) $CH_4 \rightarrow C + 4H, \Delta H = 320$

$$E_{C-H} = 90 \text{ cal}$$

$$C_2H_6 \rightarrow 2C + 6H, \Delta H = 360$$

$$\therefore 360 = E_{C-C} + 6E_{C-H}$$

$$\therefore E_{C-C} = 360 - 320 = 40 \text{ cal}.$$

10. (b) $H-H + Br-Br \rightarrow 2H-Br$

$$433 + 192 \qquad 2 \times 364$$

$$625 \qquad 728$$

Energy absorbed = Energy released

Net energy released = $728 - 625 = 103 \text{ kJ}$

i.e., $\Delta H = -103 \text{ KJ}$

Free energy and Work function

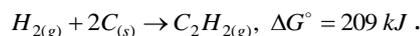
- (d) ΔG at equilibrium = 0.
- (b) For spontaneous change $\Delta G = -ve$.
- (a) When $\Delta G = -ve$ than the reaction is spontaneous in nature.
- (acd) When $\Delta H = +ve$ and $\Delta S = -ve$ than the reaction is non-spontaneous.
- (c) Because the reaction is spontaneous.
- (c) $\Delta G = -ve$.
- (a) $\Delta G_T = nRT \ln \frac{P_2}{P_1}$.
- (a) $\Delta G = 0$ for equilibrium.
- (d) At equilibrium $\Delta G = 0$.
- (a) $\Delta G^\circ = -2.303 RT \log K$
 $-4.606 = -2.303 \times 0.002 \times 500 \log K$
 $\log K = 2, K = 100$.
- (d) Spontaneous change shows $\Delta G = -ve$.
- (b) $\Delta G = \Delta H - T\Delta S, T = 25 + 273 = 298 \text{ K}$
 $= -11.7 \times 10^3 - 298 \times (-105) = 19590 \text{ J} = 19.59 \text{ kJ}$

18. (c) If $\Delta G = -ve$ reaction is spontaneous.

20. (b) $(dS)_{V,E} > 0, (dG)_{T,P} < 0$.

21. (b) Only high pressure favours the conversion.

22. (d) By $2 \times$ (ii) - (i) - (iii)



23. (b) $\Delta G = -2.303 RT \log K'$, Here $R = 2 \text{ cal}, T = 300 \text{ K}$

$$K' = \frac{10 \times 15}{3 \times 5} = 10; \Delta G = -2.303 \times 2 \times 300 \times \log_{10} 10$$

$$= -2.303 \times 2 \times 300 \times 1 = -1381.8 \text{ cal}$$

24. (a) $\Delta G = \Delta H - T\Delta S$ (Gibb's free energy equation)

25. (d) $\Delta G = \Delta H - T\Delta S$

$$\Delta G = -382.64 - (-145.6) \times 10^{-3} \times 298$$

$$= -339.3 \text{ KJ mol}^{-1}$$

26. (d) $\Delta G = \Delta H - T\Delta S; \Delta G = \Delta E + P\Delta V - T\Delta S$

For spontaneity ($\Delta G = -ve$)

$$\therefore \Delta G = -ve, \Delta E = +ve$$

27. (a) $\Delta G = \Delta H - T\Delta S = 31400 - 1273 \times 32$

$$= 31400 - 40736 = -9336 \text{ cal}$$

28. (d) $\Delta G = \Delta H - T\Delta S$

$$0 = +30.558 - T \times 0.066$$

$$\text{or } T = \frac{30.558}{0.066} = 463 \text{ K}$$

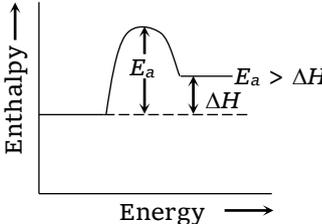
If $(dG)_{T,P} = 0$ sign '=' mean. If is reversible process

Critical Thinking Questions

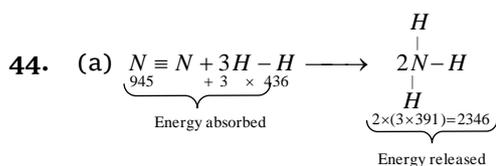
- (b) Due to randomness of particles is reduced since entropy decreases.
- (c) $\Delta H = nC_p \Delta T$
 The process is isothermal therefore $\Delta G = 0; \therefore \Delta H = 0$
- (d) $13.7 \text{ kcal} = 57 \text{ kJ} = 5.7 \times 10^4 \text{ J}$.
- (b) When strong acid and strong base neutralize each other than the value of heat generated is about 13.7 kcal .
- (b) Due to fall in temperature.
- (d) $q_p = \Delta H$.
- (d) $w = 2.303 RT \log \frac{V_2}{V_1}$. As it involves the ratio $\frac{V_2}{V_1}$, volume can be used in any units. (R should be in joules).
- (b) $-\frac{dQ}{dW} = \frac{dQ}{dQ - dE}; dE = dW + dQ; -dW = dQ - dE$
 $= \frac{nc_p dT}{nc_p dT - nc_v dT} = \frac{c_p}{(c_p - c_v)}$

450 Thermodynamics and Thermochemistry

$$= \frac{7R}{2R} \left\{ \text{for diatomic of gas } c_p = \frac{7R}{2} \right\}$$

9. (a) $C_2H_2(g) + H_2(g) \rightarrow C_2H_4(g)$
 $\Delta n_g = 1 - 2 = -1$; $\Delta H = -0.31 \text{ KJ mol}^{-1}$
 $P = 1.5 \text{ atm}$, $\Delta V = -50 \text{ mL} = -0.050 \text{ L}$
 $\Delta H = \Delta E + P\Delta V$
 $-0.31 = \Delta E - 0.0076$; $\Delta E = -0.3024 \text{ KJ}$
10. (a) Since process is exothermic then heat is evolved due to this temperature of water increases.
11. (a) Because randomness is decreases.
12. (c) Conversion of graphite into diamond is an endothermic reaction. So, heat of diamond is higher than that of graphite. But ΔS would be negative for the conversion of graphite into diamond.
13. (c) ΔH and ΔS both are +ve for spontaneous change, and $\Delta H = +ve$ for endothermic reaction.
14. (a) $\Delta G = \Delta H - T\Delta S$ is negative for spontaneity.
15. (a) Entropy (a measure of disorder) of universe is increasing toward maximum. This is II law of thermodynamics.
16. (a) For a pure substance T_A and T_B represent the same temperature. Hence A is a correct choice.
17. (a) For endothermic reaction enthalpy of products (H_p) > enthalpy of reactant (H_R)
Hence, change in enthalpy,
 $\Delta H = H_p - H_R = \text{Positive}$
18. (d) Suppose the no. of gm, 1gm for each case so the max. no of mole expected for $SO_2(g)$ which has to expected maximum entropy.
19. (b) $2Al + \frac{3}{2}O_2 \rightarrow Al_2O_3$, $\Delta H = -1596 \text{ kJ}$ (i)
 $2Cr + \frac{3}{2}O_2 \rightarrow Cr_2O_3$, $\Delta H = -1134 \text{ kJ}$ (ii)
By (i) - (ii)
 $2Al + Cr_2O_3 \rightarrow 2Cr + Al_2O_3$, $\Delta H = -462 \text{ kJ}$.
20. (a) Bomb calorimeter is used to measure heat of reaction.
21. (b) Larger the number of C and H atoms, greater is the heat of combustion.
22. (c) The bomb calorimeter is used to measure ΔE .
23. (c) 
24. (b) $\Delta H_{\text{Neutralization}} = -57.1 \text{ kJ} = -13.7 \text{ kcal}$.
25. (b) Backmann thermometer is used to measure low temperature.
26. (b) Heat required to rise the temperature of a body by 1K called thermal capacity of the body.
27. (e) Mechanical work is important only in gases as they undergo appreciable change in volume.
28. (d) It is Lavoisier and Laplace law. This is another law of thermochemistry which was put forward before Hess's law.
29. (a) It is also known as "constant heat summation" law. It state that the total amount of heat evolved or absorbed in a reaction is same whether the reaction takes place in one step or in a number of steps.
30. (c) In this reaction $\Delta n = 2 - 4 = -2$ so $\Delta H \neq \Delta E$.
31. (a) $Q = \frac{18.94 \times 0.632 \times 0.998 \times 122}{1.89} = 771.1 \text{ Kcal}$
32. (c) $\Delta H = E_a$ for forward reaction $-E_a$ for backward reaction $= 19 - 9 = 10 \text{ kJ}$.
33. (d) $Q = 18.94 \times 0.632 \times 0.998 \times 1000$
 $-57.4 + x = -12.13$
 $x = 45.2$
34. (c) Thermodynamics is a reversible process in which surroundings are always equilibrium with system.
35. (a) $1 \text{ cal} = 4.18 \text{ J} = 4.18 \times 10^7 \text{ erg} = \frac{4.18}{1.602} \times 10^{19} \text{ eV}$.
36. (b) Order of bond energy : $Cl_2 > Br_2 > F_2 > I_2$.
37. (c) Heat energy is also involved when one allotropic form of an element is converted in to another. graphite is the stabler allotrope because the heat of transformation of $C_{(\text{diamond})} \rightarrow C_{(\text{graphite})}$.
(i) $C_{(\text{dia})} + O_{2(g)} = CO_{2(g)} \Delta H = -94.5 \text{ kcal}$
(ii) $C_{(\text{graphite})} + O_{2(g)} = CO_{2(g)} \Delta H = -94.0 \text{ kcal}$
 $\Delta H_{\text{transformation}} = -94.5 - (-94.0)$
 $= -0.5 \text{ kcal}$.

38. (b) The compound Y possess less energy than the X and thus Y is more stable than X.
39. (c) Both ΔH and Δp are positive.
40. (a) For equilibrium $\Delta G^\circ = 0$.
41. (b) Equilibrium constant is decreases with temperature and ΔH is also decreases so it is -ve.
42. (a) The fusion temperature of naphthalene is minimum, because it is non-polar covalent compound and has less fusion temperature.
43. (a) The bond energy of C-H bond is $y \text{ kcal mol}^{-1}$.



$$\text{Net. energy released} = 2346 - 2253 = 93 \text{ kJ}$$

$$\text{i.e. } \Delta H = -93 \text{ kJ}.$$

45. (d) If $E = +ve$ than the cell reaction will be spontaneous.
46. (d) All are the correct statements.
47. (d) $CH_2 = CH_{2(g)} + H_{2(g)} \rightarrow H_3C - CH_{3(g)}$
- $$\begin{array}{r} 414 \times 4 = 1656 \\ 615 \times 1 = 615 \\ 435 \times 1 = \frac{435}{2706} \\ \hline \Delta H = 2706 - 2831 = -125 \text{ kJ} \end{array} \quad \begin{array}{r} 414 \times 6 = 2484 \\ 347 \times 1 = \frac{347}{2831} \\ \hline \end{array}$$
48. (b) It means that the heat absorbed when one gram molecule of HCl is formed from its elements at 298 K is
- $$= 22.060 \text{ kcal}.$$
49. (d) eq. (i) + eq. (ii) find the required result and divide by 2.
50. (d) For complete neutralization of strong acid and strong base energy released is 57.32 KJ/mol

$$\text{No. of mole of } H_2SO_4 = \frac{0.2 \times 50}{1000} = 10^{-2}$$

$$\text{No. of mole of } KOH = \frac{1}{1000} \times 50 = 5 \times 10^{-2}$$

$$So = 57.32 \times 10^{-2} = 0.5732 \text{ KJ} = 573.2 \text{ Joule}.$$

51. (b) For Combustion reaction, ΔH is negative, $\Delta n = (16 + 18) - (25 + 2) = +7$, so ΔS is +ve, reaction is spontaneous, hence ΔG is -ve.

52. (a) $\Delta G = \Delta H - T\Delta S$, $T = 27 + 273 = 300 \text{ K}$

$$\Delta G = (-285.8) - (300)(-0.163) = -236.9 \text{ kJ mol}^{-1}$$

Assertion & Reason

1. (b) Assertion and reason show, reaction at equilibrium state in which $\Delta G = 0$, $\Delta S = 0$, $\Delta H = 0$.
2. (b) It is correct that on touching the ice we feel cold because ice absorb heat from our hand.
3. (b) Entropy of ice is less than water because water molecules in solid state lose kinetic energy and hence their tendency of movement minimise. Hence entropy decrease in solid state. The reason that ice have cage like structure is also correct but reason is not or correct explanation for assertion.
4. (c) During isothermal expansion of an ideal gas against vacuum is zero because expansion is isothermal. The reason, that volume occupied by the molecules of an ideal gas is zero, is false.
5. (a) it is fact that absolute values of internal energy of substances can not be determined. It is also true that to determine exact values of constituent energies of the substance is impossible.
6. (b) Mass and volume are extensive properties. mass/volume is also an extensive parameter. Here, both assertion and reason are true.
7. (b) The molar entropy of vaporization of water it differ from ethanol due to hydrogen bonding according to VSEPR theory water molecule having two lone pair of electron by which it angular and show some polarity which is higher than that of ethanol so both assertion and reason are correct but reason is not explanation assertion.
8. (a) The assertion that the increase in internal energy for the vaporisation of one mole of water at 1 atm and 373 K is zero is true and this is because that for all isothermal process internal energy is zero.
9. (b) $N_2(g) + O_2(g) \rightleftharpoons 2NO(g)$
- We know that $\Delta H = \Delta E + \Delta nRT$
- $$\Delta n = 2 - 2 = 0 \text{ hence, } \Delta nRT = 0$$
- Therefore, $\Delta H = \Delta E$
10. (a) The enthalpies of neutralisation of strong acids and strong bases are same they are 13.7

452 Thermodynamics and Thermochemistry

Kcal. The reason is that it is heat of formation of water from H^+ and OH^- ions. $H^+ + OH^- \rightarrow H_2O, \Delta H = 13.7 \text{ Kcal}$. Thus, both assertion and reason are true.

11. (a) Zeroth (law of temperature) can also be summarized as two objects at different temperature in thermal contact with each other tend to move towards the same temperature.
12. (d) Photo chemical reactions have $+ve \Delta G$.

If a system undergoes a change in which internal energy of the system remains constant (i.e., $\Delta E = 0$), then $-W = Q$.

This means that work done by the system equals the heat absorbed by the system.

20. (e) If a refrigerator's door is kept open room gets heated as heat energy is released to the surroundings.
21. (c) Enthalpy is zero but entropy is not zero. Vibrational motion exists even at absolute zero.
22. (c) It may involve increase or decrease in temperature of the system. Systems in which such process occur, are thermally insulated from the surroundings.

14. (e) Heat absorbed in a reaction at constant temperature and constant volume $(Q_v) = E$.
15. (c) Values of state functions depend only on the state of the system and not on how it is reached.
16. (a) The properties whose magnitude depends upon the quantity of matter present in the system are called extensive properties.
17. (c) Combustion reactions are always accompanied by the evolution of heat therefore, for such reactions the value of ΔH is always negative.
18. (b) In an isothermal process change in internal energy (ΔE) is zero (as it is a function of temperature). \therefore According to first law of thermodynamics

$$\therefore q + w = \Delta E, \text{ hence } Q = -W \text{ (if } \Delta E = 0)$$